

Chemistry Exam 1 Study Guide (Chapters 1-4)

Chapter 1: Matter and Energy

1.1 Matter and its Classification:

- **Matter**- anything that occupies space and has mass.
- **Mass**- a measure of the quantity of matter.
- **Pure Substance**- matter that has the same chemical composition no matter what its origin.
 - **Elements and Compounds**
- **Mixture**- a combination of two or more substances that can vary in composition.
 - Separation of mixtures is **physical** using procedures such as **grinding, dissolving, or filtering**.
- **Homogeneous Mixture**- a combination of two or more substances that has uniform composition.
 - **Solution**- a homogeneous mixture of two or more substances uniformly dispersed at a molecular or ionic level.
- **Heterogeneous Mixture**- a combination of two or more substances that is not uniform throughout.
- **Element**- a substance that cannot be broken down into simpler substances even by a chemical reaction.
 - **Metal**- an element characterized by luster and the ability to conduct electricity. Ex: **Copper(Cu), Aluminum(Al), Iron(Fe)**
 - **Nonmetal**- an element that usually has a dull appearance and is a poor conductor of electricity. Ex: **Carbon(C), Chlorine(Cl), Sulfur(S)**
- **Compound**- a substance composed of two or more elements combined in definite proportions.
- **Chemical Formula**- describes the composition of a compound, using the symbols for the elements that make up the compound.
- **Atom**- the smallest unit of an element that has the chemical properties of that element.
- **Molecule**- two or more atoms bound together in a discrete arrangement.
- **Aqueous Solution**- a homogeneous mixture of two or more substances in which the solvent is water.

Physical States of Matter

Solid	Liquid	Gas
fixed shape	shape of container (may or may not fill it)	shape of container (fills it)
its own volume	its own volume	volume of container
no volume change under pressure	slight volume change under pressure	large volume change under pressure
particles are fixed in place and tend to be in a regular (crystalline) array	particles are randomly arranged and free to move about until they bump into one another	particles are widely separated and move independently of one another

Study Guide Introduction To Chemistry Answer Key

Iris Stovall, Donald J. DeCoste



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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text Class 8-12 Chemistry Questions and Answers PDF Arshad Iqbal, The Class 8 12 Chemistry Quiz Questions and Answers PDF Grade 8 12 Chemistry Competitive Exam Questions Chapter 1 15 Practice Tests Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions Class 8 12 Chemistry Questions and Answers PDF book covers basic concepts and analytical assessment tests Class 8 12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes The Grade 8 12 Chemistry Quiz Questions and Answers PDF eBook includes Practice material with verbal quantitative and analytical past papers questions Class 8 12 Chemistry Questions and Answers PDF Free download chapter 1 a book to review textbook questions on chapters Molecular structure acids and bases atomic structure bonding chemical equations descriptive chemistry equilibrium systems gases laboratory liquids and solids mole concept oxidation reduction rates of reactions solutions thermochemistry Questions for high school and college revision questions Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The Grade 8 12 Chemistry Interview Questions Chapter 1 15 PDF book includes high school workbook questions to practice Questions for exam Chemistry Practice Tests a textbook s revision guide with chapters Questions for NEET MCAT GRE GMAT SAT ACT competitive exam Grade 8 12 Chemistry Questions Bank Chapter 1 15 PDF book covers problem solving exam tests from chemistry practical and textbook s chapters as Chapter 1 Molecular Structure Questions Chapter 2 Acids and Bases Questions Chapter 3 Atomic Structure Questions Chapter 4 Bonding Questions Chapter 5 Chemical Equations Questions Chapter 6 Descriptive Chemistry Questions Chapter 7 Equilibrium Systems Questions Chapter 8 Gases Questions Chapter 9 Laboratory Questions Chapter 10 Liquids and Solids Questions Chapter 11 Mole Concept Questions Chapter 12 Oxidation Reduction Questions Chapter 13 Rates of Reactions Questions Chapter 14 Solutions Questions Chapter 15 Thermochemistry Questions The Molecular Structure Quiz Questions PDF e Book Chapter 1 interview questions and answers on polarity three dimensional molecular shapes The Acids and Bases Quiz Questions PDF e Book Chapter 2 interview questions and answers on Arrhenius concept Bronsted lowry concept indicators introduction Lewis concept pH strong and weak acids and bases The Atomic Structure Quiz Questions PDF e Book Chapter 3 interview questions and answers on electron configurations experimental evidence of atomic structure periodic trends quantum numbers and energy levels The Bonding Quiz Questions PDF e Book Chapter 4 interview questions and answers on ionic bond covalent bond dipole dipole forces hydrogen bonding intermolecular forces London dispersion forces metallic bond The Chemical Equations Quiz Questions PDF e Book Chapter 5 interview questions and answers on

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