

SOLUBILITY AND SOLUBILITY PRODUCT



Solubility and Solubility Product

Solubility vs Solubility Product	
DEFINITION	Solubility Property of a substance to dissolve in a solvent.
THEORY	Solubility product Mathematical description of the equilibrium between dissolved and undissolved species of a substance in solution.
FACTORS	Solubility product Describes the dissolution of a substance in a solvent.

Solubility expresses the maximum concentration of solute in a solution. The solution with the maximum solute concentration is called a saturated solution. In a saturated solution, an equilibrium occurs so that it has an equilibrium constant called the solubility product constant (K_{sp}).

- THE **SOLUBILITY PRODUCT (K_{sp})** is the product of the concentrations of ions in a saturated solution raised to the power of the coefficient at a specific temperature. For example, in the electrolyte compound $L_xZ_y(s)$, the dissociation equation for the saturated solution is as follows.



$$K_{sp} = [L^{x+}]^a [Z^{x-}]^b$$

K_{sp} RELATIONSHIP WITH SOLUBILITY (S)

- If the K_{sp} value is large, then the solubility value (s) is also large \rightarrow the substance dissolves easily
- If the K_{sp} value is small, then the solubility value (s) is also small \rightarrow the substance is difficult to dissolve

$$K_{sp} = (n - 1)^{n - 1} \times S^n \quad \text{or} \quad K_{sp} = ((a+b) - 1)^{(a+b) - 1} \times S^{(a+b)}$$

Study Guide Solubility And Concentration

Susan McMurry

Study Guide Solubility And Concentration:

Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics. Emphasizes descriptive chemistry, chemical reactions and properties while maintaining a solid treatment of chemical principles. Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises. Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered. Chapters include frequent margin comments, figures and photographs.

REHS/RS Study Guide: A Guide for Environmental Health

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993

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Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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