

# Stoichiometry

## Section 11.1 Defining Stoichiometry

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### Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule  $\text{N}_2$  + 3 molecules  $\text{H}_2$

$\text{H}_2$   $\rightarrow$  2 molecules  $\text{NH}_3$

1 mole  $\text{N}_2$  + 3 moles  $\text{H}_2$   $\rightarrow$  2 moles  $\text{NH}_3$

Mass:

$$\text{N}_2: (2 \text{ mol})(14.007 \text{ g/mol}) = 28.014 \text{ g}$$

$$3\text{H}_2: (6 \text{ mol})(1.008 \text{ g/mol}) = 6.048 \text{ g}$$

$$2\text{NH}_3: (2 \text{ mol})(14.007 \text{ g/mol}) +$$

$$(6 \text{ mol})(1.008 \text{ g/mol}) = 34.062 \text{ g}$$

$$28.014 \text{ g } \text{N}_2 + 6.048 \text{ g } \text{H}_2 \rightarrow 34.062 \text{ g } \text{NH}_3$$

34.062 g reactants = 34.062 g products



1 molecule  $\text{HCl}$  + 1 formula unit  $\text{KOH}$   $\rightarrow$

1 formula unit  $\text{KCl}$  + 1 molecule  $\text{H}_2\text{O}$

1 mole  $\text{HCl}$  + 1 mole  $\text{KOH}$   $\rightarrow$

1 mole  $\text{KCl}$  + 1 mole  $\text{H}_2\text{O}$

Mass:

$$\text{HCl: } (1 \text{ mol})(36.461 \text{ g/mol}) + (1 \text{ mol})(35.453 \text{ g/mol}) = 36.461 \text{ g}$$

$$\text{KOH: } (1 \text{ mol})(39.098 \text{ g/mol}) + (1 \text{ mol})(15.999 \text{ g/mol}) + (1 \text{ mol})(1.008 \text{ g/mol}) = 56.105 \text{ g}$$

$$\text{KCl: } (1 \text{ mol})(39.098 \text{ g/mol}) + (1 \text{ mol})(35.453 \text{ g/mol}) = 74.551 \text{ g}$$

$$\text{H}_2\text{O: } (2 \text{ mol})(1.008 \text{ g/mol}) + (1 \text{ mol})(15.999 \text{ g/mol}) = 18.015 \text{ g}$$

$$36.461 \text{ g } \text{HCl} + 56.105 \text{ g } \text{KOH} \rightarrow 74.551 \text{ g } \text{KCl} + 18.015 \text{ g } \text{H}_2\text{O}$$

92.566 g reactants = 92.566 g products



2 atoms Mg + 1 molecule  $\text{O}_2$

$\text{O}_2 \rightarrow$  2 formula units  $\text{MgO}$

2 moles Mg + 1 mole  $\text{O}_2 \rightarrow$  2 moles  $\text{MgO}$

Mass:

$$2\text{Mg: } (2 \text{ mol})(24.305 \text{ g/mol}) = 48.610 \text{ g}$$

$$\text{O}_2: (2 \text{ mol})(15.999 \text{ g/mol}) = 31.998 \text{ g}$$

$$2\text{MgO: } (2 \text{ mol})(24.305 \text{ g/mol}) +$$

$$(2 \text{ mol})(15.999 \text{ g/mol}) = 80.608 \text{ g}$$

$$48.610 \text{ g } \text{Mg} + 31.998 \text{ g } \text{O}_2 \rightarrow 80.608 \text{ g } \text{MgO}$$

80.608 g reactants = 80.608 g products

2. **Challenge:** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms Na + 2 molecules  $\text{H}_2\text{O} \rightarrow$

2 formula units  $\text{NaOH}$  + 1 molecule  $\text{H}_2$

2 mol Na + 2 mol  $\text{H}_2\text{O} \rightarrow$

2 mol  $\text{NaOH}$  + 1 mol  $\text{H}_2$

Mass:

$$2\text{Na: } (2 \text{ mol})(22.990 \text{ g/mol}) = 45.980 \text{ g}$$

$$2\text{H}_2\text{O: } (4 \text{ mol})(1.008 \text{ g/mol}) + (2 \text{ mol})(15.999 \text{ g/mol}) = 36.030 \text{ g}$$

$$2\text{NaOH: } (2 \text{ mol})(22.990 \text{ g/mol}) + (2 \text{ mol})(15.999 \text{ g/mol}) + (2 \text{ mol})(1.008 \text{ g/mol}) = 79.994 \text{ g}$$

$$\text{H}_2: (2 \text{ mol})(1.008 \text{ g/mol}) = 2.016 \text{ g}$$

$$45.980 \text{ g } \text{Na} + 36.030 \text{ g } \text{H}_2\text{O} \rightarrow$$

$$79.994 \text{ g } \text{NaOH} + 2.016 \text{ g } \text{H}_2$$

82.01 g reactants = 82.01 g products

# Study Guide Stoichiometry Answer Key

**Russo Steve Silver Michael,Steve  
Russo**

## **Study Guide Stoichiometry Answer Key:**

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on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Class 11-12 Chemistry MCQ (Multiple Choice Questions)* Arshad Iqbal,2019-05-17 The Class 11 12 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF College Chemistry MCQ PDF Download Quiz Questions Chapter 1 6 Practice Tests with Answer Key 11th 12th Grade Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Class 11 12 Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Class 11 12 Chemistry MCQ PDF book helps to practice test questions from exam prep notes The Class 11 12 Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs Class 11 12 Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters atomic structure basic chemistry chemical bonding chemistry experimental techniques gases liquids and solids tests for college and university revision guide Class 11 12 Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book Grade 11 12 Chemistry MCQs Chapter 1 6 PDF includes college question papers to review practice tests for exams Class 11 12 Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for NEET MCAT GRE GMAT SAT ACT competitive exam College Chemistry Mock Tests Chapter 1 6 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Atomic Structure MCQ Chapter 2 Basic Chemistry MCQ Chapter 3 Chemical Bonding MCQ Chapter 4 Experimental Techniques MCQ Chapter 5 Gases MCQ Chapter 6 Liquids and Solids MCQ The Atomic Structure MCQ PDF e Book Chapter 1 practice test to solve MCQ questions on Atoms atomic spectrum atomic absorption spectrum atomic emission spectrum molecules azimuthal quantum number Bohr s model Bohr s atomic model defects charge to mass ratio of electron discovery of electron discovery of neutron discovery of proton dual nature of matter electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements energy of revolving electron fundamental particles Heisenberg s uncertainty principle hydrogen spectrum magnetic quantum number mass of electron metallic crystals properties Moseley law neutron properties orbital concept photons wave number Planck s quantum theory properties of cathode rays properties of positive rays quantum numbers quantum theory Rutherford model of atom shapes of orbitals spin quantum number what is spectrum x rays and atomic number The Basic Chemistry MCQ PDF e Book Chapter 2 practice test to solve MCQ questions on Basic chemistry atomic mass atoms molecules Avogadro s law combustion analysis empirical formula isotopes mass spectrometer molar volume molecular ions moles positive and negative ions relative abundance spectrometer and stoichiometry The Chemical Bonding MCQ PDF e Book Chapter 3 practice test to solve MCQ questions on Chemical bonding chemical combinations atomic radii atomic radius periodic table atomic ionic and covalent radii atoms and

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