

### Unit 1 Study Guide — ANSWER KEY

1. Define scientific method: **a procedure for collecting, testing, and communicating information**
2. Define hypothesis: **a suggested answer to the question, an educated guess**
3. Define experiment: **a scientific test so see what will happen**
4. Define dependent variable: **the variable that you do NOT control. (on a graph is usually on the Y axis)**
5. Define independent variable: **the variable that I control. (on a graph is usually on the X axis)**
6. Define control: **the part of the experiment where everything is left the same.**
7. Define observation: **information collected with any of your 5 senses**
8. Define density: **the ratio of mass per unit volume (D = m/v)**
9. Define accuracy: **the measure of how close your data is to the ACTUAL or true value.**
10. Define precision: **a measure of how close your data (measurements) are to each other.**
11. Why do scientists use SI units of measurement? **So that measurements can be understood internationally (around the world).**
12. List (in the correct order) the steps of the scientific method.

- |   |   |
|---|---|
| <ol style="list-style-type: none"> <li>1. make an observation</li> <li>2. ask a question</li> <li>3. form a hypothesis</li> </ol> | <ol style="list-style-type: none"> <li>4. test hypothesis (experiment)</li> <li>5. analyze data/form a conclusion</li> <li>6. develop a theory</li> </ol> |
|---|---|

13. What mnemonic device is used to remember the metric system?

**K                      H                      Da                      B                      d                      c                      m**  
**g, m, L.**

14. What is the sentence (mnemonic) that will help you remember the METRIC prefixes in order?

**King Henry Died By Drinking Chocolate Milk**

15. Fill in the table below describing the metric system:

	<b>Kilo</b>	<b>Hecto</b>	<b>Deca</b>	<b>Base</b>	<b>Deci</b>	<b>Centi</b>	<b>Milli</b>
Abbreviation	<b>K</b>	<b>H</b>	<b>Da</b>	<b>g, m, L</b>	<b>d</b>	<b>c</b>	<b>m</b>
Meaning in powers of 10	<b><math>10^3</math></b>	<b><math>10^2</math></b>	<b><math>10^1</math></b>	<b><math>10^0</math></b>	<b><math>10^{-1}</math></b>	<b><math>10^{-2}</math></b>	<b><math>10^{-3}</math></b>

#### Convert the following:

- |   |   |
|---|---|
| 16. 25L to <u>25000</u> mL<br>17. 2.3 km to <u>230,000</u> cm<br>18. 0.045 g to <u>45</u> mg<br>19. 0.35 kg to <u>350</u> g<br>20. 300 mL to <u>0.300</u> L | 21. 25 cm to <u>0.25</u> m<br>22. 56784 mm to <u>0.056784</u> km<br>23. 15 g to <u>0.015</u> kg<br>24. 3.452 kL to <u>3452</u> L<br>25. 47 m to <u>47000</u> mm |
|---|---|

#### Scientific Notation:

Convert the following to scientific notation:

- |  |   |
|--|---|
| 26) 0.005 <u><math>5 \times 10^{-3}</math></u><br>27) 5.050 <u><math>5.05 \times 10^1</math></u><br>28) 0.0008 <u><math>8 \times 10^{-4}</math></u><br>29) 1,000 <u><math>1 \times 10^3</math></u> | 30) 0.025 <u><math>2.5 \times 10^{-2}</math></u><br>31) 0.0025 <u><math>2.5 \times 10^{-3}</math></u><br>32) 500 <u><math>5 \times 10^2</math></u><br>33) 5,000 <u><math>5 \times 10^3</math></u> |
|--|---|

# Study Guide The Answer Key

**Patricia A. Potter, Anne Griffin  
Perry, Patricia Stockert, Amy  
Hall, Geralyn Ochs**

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## **Table of Contents Study Guide The Answer Key**

1. Understanding the eBook Study Guide The Answer Key
  - The Rise of Digital Reading Study Guide The Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide The Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide The Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide The Answer Key
  - Personalized Recommendations
  - Study Guide The Answer Key User Reviews and Ratings
  - Study Guide The Answer Key and Bestseller Lists

5. Accessing Study Guide The Answer Key Free and Paid eBooks
  - Study Guide The Answer Key Public Domain eBooks
  - Study Guide The Answer Key eBook Subscription Services
  - Study Guide The Answer Key Budget-Friendly Options
6. Navigating Study Guide The Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide The Answer Key Compatibility with Devices
  - Study Guide The Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide The Answer Key
  - Highlighting and Note-Taking Study Guide The Answer Key
  - Interactive Elements Study Guide The Answer Key
8. Staying Engaged with Study Guide The Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide The Answer Key
9. Balancing eBooks and Physical Books Study Guide The Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide The Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide The Answer Key
  - Setting Reading Goals Study Guide The Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide The Answer Key
  - Fact-Checking eBook Content of Study Guide The Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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