

## Solutions, Acids, Bases Study Guide

### Vocabulary & Topics

|   |  |
|---|--|
| Solution – solute, solvent                          | Base – strong, weak  |
| Polar molecule                                      | Indicator, universal indicator   |
| Dissociation  | Hydronium ions – $\text{H}_3\text{O}^+$                                |
| Concentration – concentrated, dilute                | Hydroxide ions – $\text{OH}^-$   |
| Solubility – unsaturated, saturated, supersaturated | pH – 1-14 (7 is neutral)   |
| Alloy   | Neutralization   |
| Acid – Strong, weak                                 | Salts – ionic compounds formed from neutralization of acids and bases. |

### Practice Questions – Answers on the website

What are three factors that affect rate of dissolution?

Solute size – decreasing the size of the individual particles or grains of solute can increase surface area. Increased surface area exposes more solute to the solvent, speeding up the dissolution.

Temperature – increased temperature speeds up the molecules of the solvent, allowing more solvent to come into contact with the solute.

Stirring or mixing – mechanically mixing the solvent and solute allows more solvent to come into contact with the solute.

What is created when an acid is combined with a base?

An acid and a base can be combined in what is known as a neutralization reaction. In this reaction, the pH will become something between the pH's of the acid and the base. When they are combined, the products are water (from the  $\text{H}^+$  and  $\text{OH}^-$  ions) and an ionic compound known as a salt.

Name 3 characteristics that acids and bases share.

Acids and bases can both be

corrosive

poisonous

electrolytes (can conduct electricity)

Both can also be described as having unique tastes and capable of changing the color of indicators.

Describe the process of dissociation of phosphoric acid ( $\text{H}_3\text{PO}_4$ ).

During the dissociation of acid, the solute ( $\text{H}_3\text{PO}_4$ ) is dissolved in water. The polar (having slightly positive and negative ends) water molecules surround the solute. The slightly negative ends of the water molecules are attracted to the positive ends of the compounds ( $\text{H}^+$ ). The slightly positive ends of the water molecules are attracted to the

# Study Guide On Bases

**Joseph Ernest McAfee**

## **Study Guide On Bases:**

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03

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**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14. Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd numbered exercises.

**The Slide Rule** Warren Clarence Young, 1962

**MCAT Complete; Medical College Admission Test Comprehensive Study Guide Physical Sciences; Biological Sciences; Verbal Reasoning** James Adams, Maryn Munley, Nathan Eisenbeis, Caroline Miller, Jesse

Shetler, 2021-10-12. The MCAT (The Medical College Admission Test) MCAT is a standardized multiple choice examination designed to assess the examinee's problem solving, critical thinking and knowledge of science concepts and principles. Prerequisite to the study of medicine. Scores are reported in Physical Sciences, Verbal Reasoning and Biological Sciences. Study guide covers AAMC (Association of American Medical Colleges) content. Physical Sciences, Biological Sciences, Verbal Reasoning, Mathematics Concepts, The Cell, Chromosomes, Reproduction, Implantation, Microorganisms, Biochemistry, Human Physiology, The Heart, The Lymphatic System, GI Tract, Musculoskeletal System, Kidney, Hormones, Nerves, Skin, Genetics, Populations and Evolution, Elements, Hund's Rule and Radiation, The Periodic Table, Covalent Bonds, Molecular Shapes, General Chemistry, Kinetic Molecular Theory, Phase Change, Solutions, Oxidation Numbers, Entropy, Acids and Bases, Galvanic and Electrolytic Cells, Carbon, Stereochemistry, Alkanes and Alkenes, Hydrogen Bonding, Alcohols, Phenols, Aldehydes and Ketones, in Carboxylic Acids, Ether, Ammonia, Amino Acids, Carbohydrates, Spectroscopy, Separation and Distillation, Vectors and Simple Motion, Forces, Circular and Projectile Motion, Statics, Center of Gravity, Work and Energy, Power and Momentum, Stress and Strain, Elasticity and Density, Hydrostatic Pressure, Fluids in Motion, Electricity and Magnetism, Plates, Capacitors, Voltage, Batteries, Resistors, Magnetic Fields, Waves and Periodic Motion, Sound Waves, Doppler Effect, Simple Harmonic Motion, Optics, Mirrors. 205 pages.

**Research in Education**, 1966

Critical Care Study Guide Gerard J. Criner, Gilbert

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Stalling,Wayne Weiten,Ronald E. Wasden,2006-02 By Richard Stalling and Ronald Wasden of Bradley University this study

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