

REDOX REACTIONS.

NOTES...

Intro

~Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^{-}$
- Sub-points:
 - Loss of hydrogen
 - Gain of oxygen
 - Gain of **electronegative** element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $\text{Cl}_2 + 2\text{e}^{-} \rightarrow 2\text{Cl}^{-}$
- Sub-points:
 - Gain of hydrogen
 - Loss of oxygen
 - Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

~REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^{-}$

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Iris Stovall, Donald J. DeCoste



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1. Understanding the eBook Study Guide Redox Reactions
 - The Rise of Digital Reading Study Guide Redox Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Redox Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Redox Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Redox Reactions
 - Personalized Recommendations
 - Study Guide Redox Reactions User Reviews and Ratings
 - Study Guide Redox Reactions and Bestseller Lists

5. Accessing Study Guide Redox Reactions Free and Paid eBooks
 - Study Guide Redox Reactions Public Domain eBooks
 - Study Guide Redox Reactions eBook Subscription Services
 - Study Guide Redox Reactions Budget-Friendly Options
6. Navigating Study Guide Redox Reactions eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Redox Reactions Compatibility with Devices
 - Study Guide Redox Reactions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Redox Reactions
 - Highlighting and Note-Taking Study Guide Redox Reactions
 - Interactive Elements Study Guide Redox Reactions
8. Staying Engaged with Study Guide Redox Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Redox Reactions
9. Balancing eBooks and Physical Books Study Guide Redox Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Redox Reactions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Redox Reactions
 - Setting Reading Goals Study Guide Redox Reactions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Redox Reactions
 - Fact-Checking eBook Content of Study Guide Redox Reactions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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