

Study Guide:

Thermochemistry

What is thermochemistry and why is it important?

Thermo chemistry is the study of heat energy transfer in the universe. It is important to understand Thermochemistry as it provides a foundation as to how to manipulate and preserve heat in systems. It also helps us to understand how energy is disbursed in the universe as well as explaining to us how every day processes occur.

ΔH = enthalpy

IMPORTANT DEFINITIONS:

- 1) Spontaneous process: a process that occurs without any outside assistance or manipulation
 - a. Ex: Gas Diffusion
- 2) Entropy: a measure of how ordered the universe or system is
- 3) Enthalpy: change in heat in a system
- 4) System: portion of the universe that is isolated for experiment
- 5) Surroundings: everything in the universe outside of the isolated system
- 6) Universe: system + surroundings

Three Laws of Thermochemistry:

1. The total energy in the universe is constant
2. For any spontaneous process, the entropy of the universe increases ($\Delta S_{univ} > 0$)
3. The entropy of a perfect crystal at absolute zero is zero
 - a. Only one possible arrangement with no molecular movement and a lower bound for entropy

Entropy:

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DP Hallahan

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