

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$

Calculate the number of moles of oxygen needed to oxidize 12.5 g of $C_6H_{12}O_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, $KClO_3$,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $PtCl_2(NH_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

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NARAYAN CHANDER

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Table of Contents Study Guide Stoichiometry Answer Key Pearson

1. Understanding the eBook Study Guide Stoichiometry Answer Key Pearson
 - The Rise of Digital Reading Study Guide Stoichiometry Answer Key Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Stoichiometry Answer Key Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Study Guide Stoichiometry Answer Key Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Stoichiometry Answer Key Pearson
 - Personalized Recommendations
 - Study Guide Stoichiometry Answer Key Pearson User Reviews and Ratings
 - Study Guide Stoichiometry Answer Key Pearson and Bestseller Lists
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 - Study Guide Stoichiometry Answer Key Pearson Enhanced eBook Features
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 - Highlighting and Note-Taking Study Guide Stoichiometry Answer Key Pearson
 - Interactive Elements Study Guide Stoichiometry Answer Key Pearson
- 8. Staying Engaged with Study Guide Stoichiometry Answer Key Pearson
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Stoichiometry Answer Key Pearson
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Stoichiometry Answer Key Pearson
 - Setting Reading Goals Study Guide Stoichiometry Answer Key Pearson
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Stoichiometry Answer Key Pearson
 - Fact-Checking eBook Content of Study Guide Stoichiometry Answer Key Pearson
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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