

# STAAR CHEMISTRY REFERENCE MATERIALS



## OTHER FORMULAS

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$D = \frac{m}{v}$$

$$\text{Percent error} = \left( \frac{\text{accepted value} - \text{experimental value}}{\text{accepted value}} \right) (100)$$

$$\text{Percent yield} = \left( \frac{\text{actual yield}}{\text{theoretical yield}} \right) (100)$$

## CONSTANTS AND CONVERSIONS

Avogadro's number =  $6.02 \times 10^{23}$  particles per mole

$h$  = Planck's constant =  $6.63 \times 10^{-34}$  J · s

$c$  = speed of light =  $3.00 \times 10^8 \frac{\text{m}}{\text{s}}$

$K_w$  = ionization constant of water =  $1.00 \times 10^{-14} \left( \frac{\text{mol}}{\text{L}} \right)^2$

alpha particle ( $\alpha$ ) =  ${}^4_2\text{He}$       beta particle ( $\beta$ ) =  ${}^0_{-1}\text{e}$       neutron =  ${}^1_0\text{n}$

standard temperature and pressure (STP) = 0°C and 1 atm

0°C = 273 K

volume of ideal gas at STP = 22.4  $\frac{\text{L}}{\text{mol}}$

1 cm<sup>3</sup> = 1 mL = 1 cc

1 atm = 760 mm Hg = 101.3 kPa

$R$  = ideal gas constant =  $0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} = 8.31 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}} = 62.4 \frac{\text{L} \cdot \text{mm Hg}}{\text{mol} \cdot \text{K}}$

1 calorie (cal) = 4.18 joules (J)

1000 calories (cal) = 1 Calorie (Cal) = 1 kilocalorie (kcal)

## RULES FOR SIGNIFICANT FIGURES

1. Non-zero digits and zeros between non-zero digits are always significant.
2. Leading zeros are not significant.
3. Zeros to the right of all non-zero digits are only significant if a decimal point is shown.
4. For values written in scientific notation, the digits in the coefficient are significant.
5. In a common logarithm, there are as many digits after the decimal point as there are significant figures in the original number.

# Staar Chemistry Study Guide

**J Spring**

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