

Study Buddy: Classification of Chemical Reactions

Terms	Definitions	Representation									
Law of Conservation of Mass	During a chemical or physical change, mass is conserved. Mass is neither created nor destroyed. Mass of reactants equals mass of products.	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ <table border="1"> <tr> <td>N_2</td> <td>3H_2</td> <td>$\rightarrow 2\text{NH}_3$</td> </tr> <tr> <td>28g</td> <td>6g</td> <td>34g</td> </tr> <tr> <td>2 atoms</td> <td>6 atoms</td> <td>8 atoms</td> </tr> </table>	N_2	3H_2	$\rightarrow 2\text{NH}_3$	28g	6g	34g	2 atoms	6 atoms	8 atoms
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28g	6g	34g									
2 atoms	6 atoms	8 atoms									
Precipitation Reaction	Occurs when two aqueous solutions react to produce a solid precipitate.	$\text{NaCl}_{(aq)} + \text{AgNO}_{(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{(aq)}$									
Redox Reaction	Occurs when one or more electrons is transferred during a chemical reaction. Reduction: occurs when a species gains an electron Oxidation: occurs when a species loses an electron	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$ $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^-$									

Representations:

1. Balance the following chemical equations and draw a particle diagram showing conservation of mass:

a. Hydrogen gas and oxygen gas react to form liquid water.



b. Solid zinc is added to an aqueous solution of silver nitrate to produce zinc nitrate and solid silver.



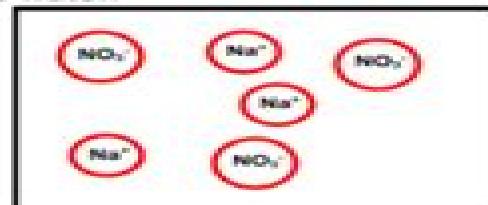
2. A chemist added a 5.0g piece of solid magnesium metal to 10.0g of hydrochloric acid ($\text{HCl}_{(aq)}$) in a 50 mL beaker. During the reaction, the solution in the beaker became warmer and produced many bubbles. Once the reaction finished, the mass of the final solution was only 14.6g. What is the best explanation for the "missing" mass?

A bubbles are an indicator that a gas is being produced. In this case, 0.4g of gas were produced and no longer in the solution.

0.21 mole of Mg added to excess HCl produces 0.21 moles of H_2 , or 0.4g of H_2 .

3. Describe how you would make 100. mL of 2.0M NaNO_3 from powdered NaNO_3 . In the box, show what happens to the NaNO_3 when it is added to water.

- Determine mass of NaNO_3 needed:
 $2.0\text{M} = (\text{X g}/58.5 \text{ g/mol})/0.100 \text{ L} \quad \text{X} = 17\text{g}$
- Put 17g of NaNO_3 in a 100 mL volumetric flask
- Add water to 100 mL
- Mix



*Water molecules not shown for clarity of ions.

Study Guide Redox Reaction

John A. Olmsted, Gregory M. Williams

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