

# Chapter 19 Oxidation-Reduction Reactions

## Chapter Opener

\_\_ **Chapter Overview, TE** Review the objectives listed in the Student Edition.

## Section 1 Oxidation and Reduction

**PACING: 45 minutes**

**PENNSYLVANIA ACADEMIC STANDARDS FOR SCIENCE AND TECHNOLOGY:**

3.1.E.1 Describe how fundamental science and technology concepts are used to solve practical problems.

## Objectives

1. Assign oxidation numbers to reactant and product species.
2. Define *oxidation* and *reduction*.
3. Explain what an oxidation-reduction (redox reaction) is.

**Focus (5 minutes)**

\_\_ **Lesson Starter, TE** Review oxidation numbers and have students familiarize themselves with the rules for assigning oxidation numbers in **Table I**.

**MOTIVATE (10 minutes)**

\_\_ **Demonstration, A Spontaneous Redox Reaction, TE** Demonstrate a redox reaction by using KI,  $\text{KMnO}_4$ , and HCl and make sure students can explain the formation of the precipitate and the color changes.

**TEACH (20 minutes)**

\_\_ **PowerPoint**

\_\_ **Visual Strategy, Figure 1, TE** Make sure students understand how this figure illustrates how an element can exist in a number of different oxidation states. (**GENERAL**)

\_\_ **Chemistry in Action, Photochromic Lenses, SE** Photochromic eyeglasses darken and become transparent according to ultraviolet light because of an oxidation-reduction reaction.

\_\_ **Application, TE** Sodium hypochlorite,  $\text{NaOCl}$ , is the chemical name for household bleach used to remove stains from clothing through a redox reaction.

\_\_ **Discussion, TE** Students should follow up the spontaneous redox reaction demonstration by explaining the results in terms of oxidation and reduction and by writing formulas, using **Table I** for the oxidation numbers.

\_\_ **Chemistry in Action, Skunk-Spray Remedy, SE** Paul Krebaum invented a better way than tomato juice to get rid of skunk spray odor.

# Study Guide Redox Reactions Answers

**O García**



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