

The Mole Study Guide

Mass Composition: What's your element's % of mass after the reaction? Molar mass to obtain your value!

10. There are 3.75 moles of CO_2 .

11. There are 3.75 moles of CO_2 .

12. There are 3.75 moles of CO_2 .

13. There are 3.75 moles of CO_2 .

14. There are 3.75 moles of CO_2 .

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20. There are 3.75 moles of CO_2 .

21. There are 3.75 moles of CO_2 .

22. There are 3.75 moles of CO_2 .

23. There are 3.75 moles of CO_2 .

24. There are 3.75 moles of CO_2 .

25. There are 3.75 moles of CO_2 .

26. There are 3.75 moles of CO_2 .

Name _____ Period _____ Date _____

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Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_5\text{H}_{10}\text{O}_5$	

Calculate the Molar Mass for the following compounds.

10. H_2O

11. H_2O_2

12. H_2O

13. H_2O_2

Calculate the Percent Composition for each element in the following compound.

10. H_2O

11. H_2O_2

12. H_2O

13. H_2O_2

14. H_2O

15. H_2O_2

16. H_2O

17. H_2O_2

18. H_2O

19. H_2O_2

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 58.80% Hg, 21.80% Cl, 1.40% N. The molar mass for Aqueous is 100.0 g/mol.

a. If you have 100 g of Aqueous, how many grams of Hg, Cl, and N do you have?

b. Determine the Empirical Formula for Aqueous.

c. Determine the Molecular Formula for Aqueous.

Study Guide The Mole

Holder



Study Guide The Mole:

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