

Gas Laws Study Guide

Write Gas Laws Principles, identify the variables, & describe the trends in the following table.

1. A sample of gas is at 1.0 atm, 273 K, and 22.4 L.

2. The volume of a gas is 1.0 L at 1.0 atm and 273 K.

3. A sample of gas is at 1.0 atm, 273 K, and 22.4 L.

4. A sample of gas is at 1.0 atm, 273 K, and 22.4 L.

5. A sample of gas is at 1.0 atm, 273 K, and 22.4 L.

6. A sample of gas is at 1.0 atm, 273 K, and 22.4 L.

Name: _____ Date: _____

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Standard Temperature = _____ K, _____ °C

Standard Pressure = _____ atm, _____ mmHg, _____ kPa, _____ psi

Standard Molar Volume = _____ L, _____ mL

Write the formula for the following:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Combined Gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Charles' Law

$$P_1 V_1 = P_2 V_2$$

Boyle's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Gay-Lussac's Law

$$PV = nRT$$

Ideal Gas Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Avogadro's Law

Complete the following statements by filling in the blanks & circling the correct words.

As Pressure increases (increases/decreases) (increases/decreases) given that _____ is constant, meaning that they have a direct/inverse relationship.

As volume increases (increases/decreases) (increases/decreases) given that _____ is constant, meaning that they have a direct/inverse relationship.

As Pressure increases (increases/decreases) (increases/decreases) given that _____ is constant, meaning that they have a direct/inverse relationship.

As the number of moles of gas in a container increases the volume increases (increases/decreases), the Pressure increases (increases/decreases), and the Temperature increases (increases/decreases).

Study Guide The Gas Laws

Steven S. Zumdahl, Paul B. Kelter



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