

THERMOCHEMICAL EQUATIONS AND TYPES OF ENTHALPY CHANGES



Thermochemical Equations

A thermochemical equation is a reaction equation accompanied by an enthalpy change (ΔH). In addition to stating the number of moles reactants and the number of moles product, it also displays the amount of heat released or absorbed in the reaction in units of kJ.

Example :

The thermochemical equation is known as:



This means that the formation of 1 mol H_2O hydrogen gas and oxygen gas released energy of 285.85 kJ (exothermic reaction).

Some things that must be considered in the thermochemical equation are

1. THE REACTION COEFFICIENT EXPRESSES MOLES.

Example:



(means that when 1 mol H_2 reacts with 1 mol I_2 to produce 2 mol HI , heat is absorbed by 51.9 kJ).

2. IF THE REACTION IS REVERSED, ΔH CHANGES SIGN

From reaction 1:



3. IF THE REACTIONS ARE MULTIPLIED, THEN ΔH IS MULTIPLIED

From reaction 1:



Study Guide Thermochemical Equations Answers

Clyde Metz

Study Guide Thermochemical Equations Answers:

Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences [Chemistry, Student Study Guide](#) James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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