

# CHEMISTRY Terminology

The ABCs of chemistry

**A**

**A:** Symbol for atomic mass number  
**absolute zero:** Lowest temperature; 0 Kelvin  
**Ac:** Symbol for the element actinium,  $Z = 89$   
**accuracy:** The closeness of measurement to true value  
**acetate anion:**  $\text{CH}_3\text{COO}^-$ , charge -1  
**acetic acid:**  $\text{C}_2\text{H}_4\text{O}_2$   
**acetone:** Dimethyl ketone; organic solvent  
**acid-base indicator:** A chemical that changes color if the pH changes  
**acid-base titration:** A method for determining acid or base concentrations  
**acid definitions:** See Arrhenius acid, Brønsted-Lowry acid, Lewis acid  
**acid ionization constant ( $K_{\text{a}}$ ):** Equilibrium constant for acid dissociation  
**acidic solution:** pH below 7  
**actinides:** Element with  $Z = 90$  or above; radioactive  
**activation energy ( $E_a$ ):** A process's energy barrier  
**affiliation:** Attraction of unlike molecules or materials  
**Ag:** Symbol for the element silver,  $Z = 47$   
**Al:** Symbol for the element aluminum,  $Z = 13$   
**alcohol:** Organic compound with -OH group;  $\text{R}-\text{OH}$   
**aldehydes:** Carbonyl group bonded to hydrogen and an organic group  
**alkali metal family:** Lithium, sodium, potassium, rubidium, cesium, and francium; column #1  
**alkaline earth metal family:** Beryllium, magnesium, calcium, strontium, barium, and radium; column #2  
**alkane:** Hydrocarbon; all C-C single bonds  
**alkene:** Hydrocarbon; 1 or more C=C double bonds  
**alkyne:** Hydrocarbon; 1 or more C≡C triple bonds  
**allotropes:** Two or more forms of an element  
**alloy:** Solution of 2 or more metals  
**alpha ( $\alpha$ ):** Greek letter that denotes radioactive particle and various scientific variables  
**alpha particle:** Helium nucleus; charge +2  
**Am:** Symbol for the element americium,  $Z = 95$   
**amalgam:** Alloy of mercury with other metal  
**amine:** Organic base;  $\text{R}_3\text{N}$ ; modified ammonia  
**amine acid:** Compound with organic acid and organic base property; found proteins and peptides  
**ammonium:**  $\text{NH}_4^+$ ; base  
**ammonium cation:**  $\text{NH}_4^+$ ; charge +1  
**ammonium hydroxide:**  $\text{NH}_4\text{OH}$   
**amphoteric oxide:** Exhibits both acid and basic properties  
**amu:** Atomic mass unit;  $1/12$  mass of  $\text{C}-12$   
**anion:** Ion with a negative charge  
**anode:** Electrode that supports oxidation  
**antibonding MO:** MO is less stable than sigma MOs  
**AO:** Abbreviation for atomic orbital



**aq:** Aquous; "dissolved in water"  
**Ar:** Symbol for the element argon,  $Z = 18$   
**aromatic:** Organic compound with a benzene ring  
**Arrhenius acid:** Produces hydronium ion in water solution  
**Arrhenius base:** Yields hydroxide ion in water solution  
**Arrhenius equation for rate constant ( $k$ ):**

$$k = A e^{-E_a/RT}$$

$E_a$  = activation energy

**arsenide anion:**  $\text{As}^{3-}$ ; charge -3  
**As:** Symbol for the element arsenic,  $Z = 33$   
**At:** Symbol for the element astatine,  $Z = 85$   
**atom:** Smallest unit in "atmosphere"  
**atom:** Fundamental unit of all matter  
**atomic mass number (A):** Total number of protons and neutrons  
**atomic number (Z):** Number of protons in the nucleus  
**atomic orbital:** Wave motion of electrons in atoms  
**atomic radius:** Empirical measure of atom size  
**atomic weight:** Weighted average of natural isotopes of an element  
**Au:** Symbol for the element gold,  $Z = 79$   
**Aufbau principle:** Order the filling of electronic subshells of the elements  
**Avogadro's law:** Volume is proportional to gas moles at fixed pressure and temperature  
**Avogadro's number:**  $N_A = 6.022 \times 10^{23}$  mol<sup>-1</sup>

**B:** Symbol for the element boron,  $Z = 5$   
**Ba:** Symbol for the element barium,  $Z = 56$   
**balanced equations:** Each side of the equation has the same number of atoms of each element and charge  
**balanced redox:** Electron loss = electron gain  
**barium cation:**  $\text{Ba}^{2+}$ ; charge +2  
**base definition:** See Arrhenius base, Brønsted-Lowry base, Lewis base  
**base ionization constant ( $K_b$ ):** Equilibrium constant for base dissociation  
**basic solution:** pH over 7  
**battery:** Source of electrical power; galvanic cell  
**Ber:** Symbol for the element beryllium,  $Z = 4$   
**benzene:**  $\text{C}_6\text{H}_6$ ; organic solvent; aromatic ring of 6 carbon atoms  
**beryllium cation:**  $\text{Be}^{2+}$ ; charge +2  
**beta ( $\beta$ ):** Greek letter that denotes radioactive particle and a number of scientific variables  
**beta particles:** Energetic electrons from the nucleus; charge -1  
**Bh:** Symbol for the element bismuth,  $Z = 107$   
**Bk:** Symbol for the element berkelium,  $Z = 97$   
**bicarbonate anion:**  $\text{HCO}_3^-$ ; charge -1  
**bimolecular oxygen:**  $\text{O}_2$   
**binary compound:** Formed from 2 elements  
**bisulfate anion:**  $\text{HSO}_4^-$ ; charge -1



**Bk:** Symbol for the element berkelium,  $Z = 97$   
**bleach:** Chemical that is a strong oxidizing agent  
**boiling:** Liquid  $\rightarrow$  gas at the boiling point  
**boiling point elevation:** A solution has a higher boiling point than a pure solvent  
**boiling point (T<sub>b</sub>):** Liquid-gas equilibrium;  $P = 1$  atm  
**boiling point of water:** 100°C  
**bond energy:** Energy held by a chemical bond  
**bond length:** Distance between 2 atoms in a bond  
**bond order for Lewis model:** # of bonds divided by # of bonded atoms  
**bond order for MO treatment:** # of filled antibonding MOs  
**bonding electrons:** Form chemical bonds in a compound  
**bonding MO:** MO is more stable than AOs  
**Boyle's law:**  $PV =$  constant for gas; Boyle's law  
**Br:** Symbol for the element bromine,  $Z = 35$   
**bromide anion:**  $\text{Br}^-$ ; charge -1  
**brønsted-Lowry acid:** Proton donor  
**brønsted-Lowry base:** Proton acceptor  
**buffer:** Composed of weak acid and weak base; serves to keep pH constant



**C:** Symbol for carbon, the SI prefix for "10<sup>3</sup>"  
**c:** Symbol for the speed of light  
**°C:** Celsius temperature scale  
**C:** Symbol for the element carbon,  $Z = 6$   
**C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>:** Acetic acid  
**Ca:** Symbol for the element calcium,  $Z = 20$   
**calcium cation:**  $\text{Ca}^{2+}$ ; charge +2  
**calcium fluoride:**  $\text{CaF}_2$   
**calcium hydroxide:**  $\text{Ca}(\text{OH})_2$   
**calories:** English unit of heat energy  
**calorimetry:** Study of the release or absorption of heat  
**carbide anion:**  $\text{C}^-$ ; charge -4  
**carboxylates:** Organic compound; carbon bonded to several -COO groups (sugar and starch)  
**carbon dioxide:**  $\text{CO}_2$   
**carbon monoxide:**  $\text{CO}$   
**carbonate anion:**  $\text{CO}_3^{2-}$ ; charge -2  
**carbonic acid:**  $\text{H}_2\text{CO}_3$   
**carboxyl group:**  $\text{COOH}$ ; found in carboxylic acids  
**carboxylic acid:** Organic acid;  $\text{R}-\text{COOH}$   
**catalyst:** Accelerates a reaction but is not consumed in the reaction; lowers the activation energy  
**cathode:** Electrode that supports reduction  
**cation:** Ion with a positive charge  
**Cd:** Symbol for the element cadmium,  $Z = 48$



# Study Guide Modern Chemistry

**Mary Burnham, Carol Hurd**

## Study Guide Modern Chemistry:

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