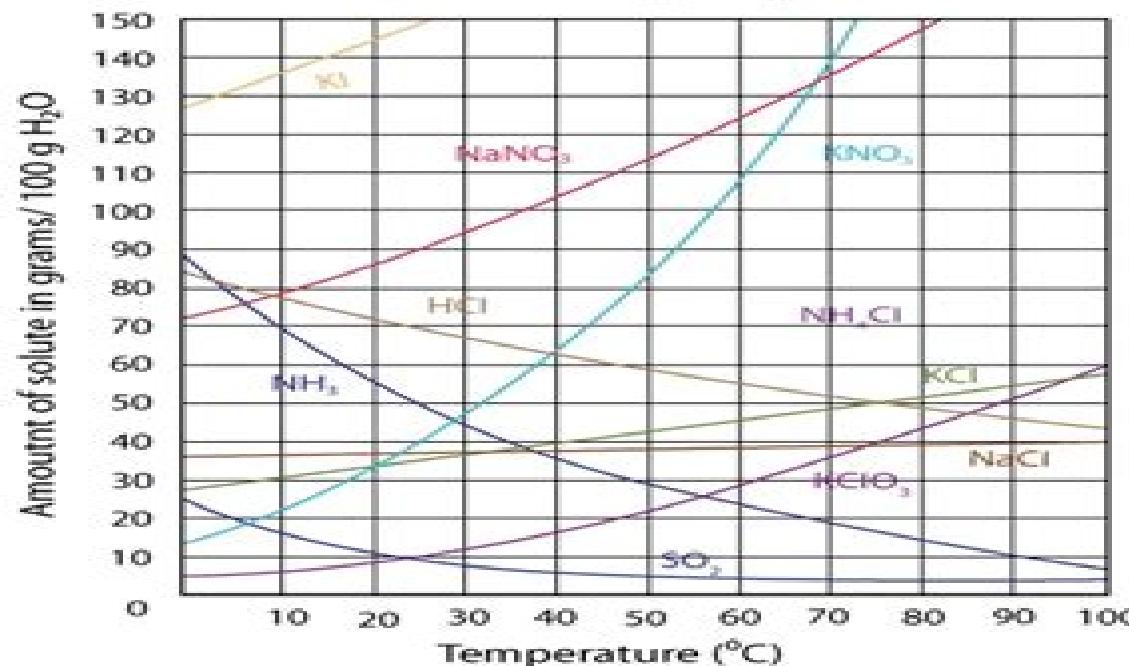


# Interpreting Solubility Curve

1. Define solubility. \_\_\_\_\_

2. Using the solubility curve below, answer the given questions.



I. What is the solvent in this solubility curve? \_\_\_\_\_

II. Saturation points are given per how much of the solvent? \_\_\_\_\_

III. What is the saturation point of  $\text{NaNO}_3$  at  $10^\circ\text{C}$ ,  $40^\circ\text{C}$ , &  $80^\circ\text{C}$ ?

$10^\circ\text{C} =$  \_\_\_\_\_       $40^\circ\text{C} =$  \_\_\_\_\_       $80^\circ\text{C} =$  \_\_\_\_\_

IV. How much  $\text{NH}_3$  can you dissolve in water at  $10^\circ\text{C}$ ,  $30^\circ\text{C}$ , &  $90^\circ\text{C}$ ?

$10^\circ\text{C} =$  \_\_\_\_\_       $30^\circ\text{C} =$  \_\_\_\_\_       $90^\circ\text{C} =$  \_\_\_\_\_

V. Which salt is the least soluble in water? \_\_\_\_\_

VI. How many grams of potassium chloride can be dissolved in 200 g of water at  $80^\circ\text{C}$ ? \_\_\_\_\_

VII. At  $40^\circ\text{C}$ , how much potassium nitrate will be dissolved in 300 g of water? \_\_\_\_\_

VIII. Which salt shows the least change in solubility from  $0^\circ$  to  $100^\circ\text{C}$ ? \_\_\_\_\_

IX. Which substance shows a decrease in solubility from  $0^\circ$  to  $100^\circ\text{C}$ ? \_\_\_\_\_

X. At  $30^\circ\text{C}$ , 90 g of sodium nitrate is dissolved in 100 g water. Is this solution saturated, unsaturated, or supersaturated? \_\_\_\_\_

# Study Guide Solubility Graph Key

**Susan McMurry, John McMurry**

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