

Summer Packet Answers w/ explanations

- 1.) B. The electron configuration for nitrogen is $1s^2 2s^2 2p^3$. There are three electron orientations within the p orbital and by Hund's rule, electrons will fill those orientations in such a way that they remain unpaired if possible.
- 2.) C. Fluorine the further element to the right on the periodic table. It is also at the top of its group. This, however, does not explain why it is the smallest, but merely provides an answer. Fluorine is the smallest because it has the largest number of protons while still only having used the 2p subshell. All the other elements have fewer protons that exert a smaller pull on their respective protons (smaller nuclear charge).
- 3.) A. Lithium has the smallest ionization energy because its outer most electron is by itself in the 2s subshell and experiences less nuclear force than the outer electrons in the atoms with more protons. Also, it is much closer to having a filled outer electron shell than the others.
- 4.) A. In this compounds the total charge on the molecule is zero. All others have a charge.
- 5.) B. Strontium is the best choice because of its location on the periodic table. Strontium's effective nuclear charge is much greater than the other choices.
- 6.) D. Bromine is a liquid a room temperature due to its intermolecular forces.
- 7.) D. Bromine has a charge of 1- as a ion because the gain of one electron fills bromine's outermost valence shell. Bromine has a moderate electronegativity value as well and a higher electron affinity value.
- 8.) B. Strontium has 38 protons; just the way the history works out.
- 9.) A. Potassium has a larger atomic radius than calcium because calcium has one more proton than potassium. Also, calcium's one extra electron fills its outermost valence shell.
- 10.) A. Atoms with similar electron configurations react in a similar way when reacted in a similar way. We know this is not true for protons and neutrons based on the fact that ions and isotopes do not always behave like their parent elements.
- 11.) B. Both *s* and *f* represent orbital shapes (sphere and well, flower, is the best descriptor).
- 12.) E. Since calcium electrons occupy *s* and *p* orbitals they can only be excited to either *d* or *f* orbitals.
- 13.) C. This is the best answer given the choices. Reading it now the word "shape" ought to be replaced with "orientation". This question is poor. If I feel badly enough between now and the first day of school you may get first day presents.
- 14.) D. Chlorine has 5 electrons in its outermost *p* orbital.
- 15.) C. Any electron configuration that has 6 electrons in its outermost *p* orbital is most often a noble gas, an anion of a non-metal or cation of the alkaline or alkali earth metals.
- 16.) B. The angular momentum quantum number describes the shape of an orbital and is denoted *l*.
- 17.) C. Bromide carries a charge of 1-. This was explained for an earlier question.

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