

20 Oxidation-Reduction Reactions

Big idea REACTIONS

20.1 The Meaning of Oxidation and Reduction

For students using the Foundation edition, assign problems 2–6, 8, 10–16, 18, 19.

Essential Understanding Oxidation and reduction are opposite chemical processes during which electrons are shifted between reactants.

Reading Strategy

Compare and Contrast Organizing information in a table helps you compare and contrast several topics at one time. Consider comparing and contrasting oxidation and reduction and how they relate to oxygen and electrons. As you read, ask yourself, “How are they similar? How are they different?”

As you read Lesson 20.1, use the compare and contrast table below. Complete the table, comparing how oxidation and reduction relate to electrons and oxygen.

Oxidation and Reduction		
	Compare	Contrast
Oxygen	<i>If oxygen is involved in the process, a change occurs in the number of oxygen atoms.</i>	<i>During oxidation, a substance gains oxygen. During reduction, a substance loses oxygen.</i>
Electrons	<i>A change occurs in the number of electrons.</i>	<i>During oxidation, a substance loses one or more electrons. During reduction, a substance gains one or more electrons.</i>

EXTENSION Use the information about electrons in the chart to explain why oxidation and reduction always occur together.

When electrons are gained by one substance, the electrons must come from another substance, which loses the electrons.

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Kenneth L. Henold



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