

## Chapter 8 - Chemical Equations and Reactions

### **8-1 Describing Chemical Reactions**

#### I. Introduction

##### A. Reactants

1. Original substances entering into a chemical rxn

##### B. Products

1. The resulting substances from a chemical rxn

Reactants  $\rightarrow$  Products

##### C. Chemical Equation

1. Represents with symbols and formulas, the identities and relative amounts of the reactants and products in a chemical rxn

#### II. Indications of a Chemical Reaction

##### A. Evolution of Heat and Light

1. Evidence of energy being released (exothermic rxn)

##### B. Production of a Gas

1.  $\text{CO}_2$ ,  $\text{H}_2$ ,  $\text{H}_2\text{S}$  are some gases produced by chemical rxns



##### C. Formation of a Precipitate

1. Precipitate is a solid that is produced as a result of a chemical rxn in solution  
 $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{BaSO}_4(\text{s})$

#### III. Characteristics of Chemical Equations

##### A. The equation must represent known facts

1. This can be done with a word equation

"hydrogen reacts with oxygen to form water"



##### B. The equation must contain the correct formulas for reactants and products

1. This is done with a formula equation



##### C. The law of conservation of atoms must be satisfied

1. Balancing is done with coefficients - small whole numbers that appear in front of a formula



##### D. Additional symbols used in Chemical equations

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**Arthur James Wells**

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