

## Stoichiometry Study Guide KEY

### Chemistry

### RHS – Mr. Moss

1. Define the following:
  - a. *Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.*
  - b. *Mole – The SI unit used to measure the amount of a substance that contains  $6.02 \times 10^{23}$  atoms of that substance.*
  - c. *Mole Ratio-The ratio between any two substances in a balanced chemical equation.*
  - d. *Molar Mass-The mass in grams of one mole of any substance.*
2. What relationships can be determined from a balanced chemical equation? *The relationships between particles, moles, and mass for all reactants and products.*
3. Explain why mole ratios are central to stoichiometric calculations. *Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.*
4. What is the first step in all stoichiometric calculation? (Be careful.) *Write a balanced chemical equation.*
5. How is molar mass used in some stoichiometric calculations? *Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.*
6. Find the molar mass of the following:
  - a.  $\text{H}_2\text{SO}_4$  (sulfuric acid) *98.09 g/mol*
  - b.  $\text{ZnO}$  (zinc oxide) *81.39 g/mol*
  - c.  $\text{NH}_4\text{NO}_3$  (ammonium nitrate) *80.05 g/mol*
7. List the mole ratios of the following *balanced* chemical equation.
  - a.  $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$   $\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} : \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3} : \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$  (*Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.*)

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**Leo J. Malone**

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