

The Mole

Section 10.1 Measuring Matter

page 320–324

Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol-Zn} \times \frac{6.02 \times 10^{23} \text{ atoms-Zn}}{1 \text{ mol-Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H_2O).

$$11.5 \text{ mol-H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol-H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules of H}_2\text{O}$$

3. Silver nitrate (AgNO_3) is used to make several different silver halides used in photographic films. How many formula units of AgNO_3 are there in 3.25 mol AgNO_3 ?

$$3.25 \text{ mol-AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol-AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units of AgNO}_3$$

4. Challenge Calculate the number of oxygen atoms in 5.0 mol of oxygen molecules. Oxygen is a diatomic molecule, O_2 .

$$5.00 \text{ mol-O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules-O}_2}{1 \text{ mol-O}_2} \\ \times \frac{2 \text{ O atoms}}{\text{molecule-O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?

a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms-Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms-Al}} \\ = 9.55 \text{ mol Al}$$

b. 2.50×10^{20} atoms Fe

$$2.50 \times 10^{20} \text{ atoms-Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms-Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. Challenge Identify the representative particle for each formula and convert the given number of representative particles to moles.

a. $3.75 \times 10^{24} \text{ CO}_2$

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules-CO}_2 \\ \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules-CO}_2} = 6.23 \text{ mol CO}_2$$

b. $3.58 \times 10^{23} \text{ ZnCl}_2$

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula-units-ZnCl}_2 \\ \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula-units-ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 10.1 Assessment

page 324

7. Explain why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. State the mathematical relationship between Avogadro's number and 1 mol.

One mole contains 6.02×10^{23} representative particles.

9. List the conversion factors used to convert between particles and moles.

$$\frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mol}} \\ \frac{1 \text{ mol}}{1 \text{ mol}} \\ \frac{6.02 \times 10^{23} \text{ representative particles}}{6.02 \times 10^{23} \text{ atoms}}$$

Study Guide The Mole Answer Key

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Table of Contents Study Guide The Mole Answer Key

1. Understanding the eBook Study Guide The Mole Answer Key
 - The Rise of Digital Reading Study Guide The Mole Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide The Mole Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide The Mole Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide The Mole Answer Key
 - Personalized Recommendations
 - Study Guide The Mole Answer Key User Reviews and Ratings
 - Study Guide The Mole Answer Key and Bestseller Lists
5. Accessing Study Guide The Mole Answer Key Free and Paid eBooks
 - Study Guide The Mole Answer Key Public Domain eBooks
 - Study Guide The Mole Answer Key eBook Subscription Services
 - Study Guide The Mole Answer Key Budget-Friendly Options

6. Navigating Study Guide The Mole Answer Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - Study Guide The Mole Answer Key Compatibility with Devices
 - Study Guide The Mole Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide The Mole Answer Key
 - Highlighting and Note-Taking Study Guide The Mole Answer Key
 - Interactive Elements Study Guide The Mole Answer Key
8. Staying Engaged with Study Guide The Mole Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide The Mole Answer Key
9. Balancing eBooks and Physical Books Study Guide The Mole Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide The Mole Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide The Mole Answer Key
 - Setting Reading Goals Study Guide The Mole Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide The Mole Answer Key
 - Fact-Checking eBook Content of Study Guide The Mole Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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