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 $3\text{CO} + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 2\text{H}_2\text{O}$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g $\text{Al} \times 1 \text{ mol} / 2 \text{ mol} = 0.5 \text{ mol}$ $\text{Al}_2\text{O}_3 \times 101.964 \text{ g/mol} = 50.982 \text{ g}$ $\text{Al} \times 4 \text{ mol} / 1 \text{ mol} = 4 \text{ mol}$ $\text{Al}_2\text{O}_3 \times 101.964 \text{ g/mol} = 407.856 \text{ g}$ c determine the percent yield yield actual $\times 100 = 260.0 \text{ g} / 407.856 \text{ g} \times 100 = 64.01\%$ theoretical 302.3 g

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