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# HOW TO EAT A 30G PROTEIN BREAKFAST EVERY DAY

*(with Recipes!)*

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# Protein Breakfast Latest Tutorial

**Avery Nolan**

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web start with 8.3 mole  $h_2s$  and multiply it with 2 mole  $h_2o$  over 2 mole  $h_2s$  from the equation  $8.3 \text{ mole } h_2s \times 2 \text{ mole } h_2o / 2 \text{ mole } h_2s \times 18 \text{ g } h_2o = 149.4 \text{ g } h_2o$  is your theoretical yield so percent yield  $137.1 \text{ g } h_2o / 149.4 \text{ g } h_2o \times 100 = 91.77\%$  problems 6-8 are similar to problem 5

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected × 100 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction al + o<sub>2</sub> → al<sub>2</sub>o<sub>3</sub> a) determine the theoretical yield of al<sub>2</sub>o<sub>3</sub> 160.0 g al x 1 mol al x 2 mol al<sub>2</sub>o<sub>3</sub> x 101.964 g al<sub>2</sub>o<sub>3</sub> 302.3 g al<sub>2</sub>o<sub>3</sub> 26.982 g al x 4 mol al x 1 mol al<sub>2</sub>o<sub>3</sub> b) determine the percent yield yield actual x 100 260.0 g x 100 86.01 theoretical 302.3 g

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ch3oh can be produced by the following reaction  $2\text{h}_2 + \text{co} \rightarrow \text{ch}_3\text{oh}$  calculate the theoretical yield of ch3oh if 68.5 g of co is reacted with 8.6 g of h2

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