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Personal Computing ,1979-07

Biennial Report Washington (State). Adjutant-General's Office, 1922

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Business Week, 2006

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web calorimetry problems calorimetry problems $q_{\text{sur}} = m \times c \times \Delta t$ q heat m mass $q_{\text{rxn}} = q_{\text{sur}} = t_f - t_i$ name per date c specific heat for water 4.184 J/gC 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased by 8.1 °C when 207 J of heat is added 2

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25.0 °C and heat it to 145.0 °C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 °C

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web sample problem calorimetry and enthalpy changes in an experiment 25.0 mL of 1.00 M HCl at 25.0 °C is added to 25.0 mL of 1.00 M NaOH at 25.0 °C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 °C calculate the enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/mL and that

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chemical mechanical or electrical reaction it also helps to measure the heat capacity of various

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web calorimetry worksheet 1 if 0 315 moles of hexane C_6H_{14} is combusted in a bomb calorimeter containing 5 65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55 4 C the specific heat capacity of water is 4 184 J g C h ms t h 5 650 grams H_2O 4 184 J g C 55 4 C h 1310 kJ

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs $1.41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6 23 $\times 10^4 \text{ J}$ 3 1 71 $\times 10^4 \text{ J}$ 4 1 11 $\times 10^5 \text{ J}$ 5 2 11 $\times 10^5 \text{ J}$ 6

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