

Matter and Chemical Change Topic 1-5

1. Explain or describe how we use composition of matter to organize the world around us. Provide examples of pure substances and mixtures.

Pure substance:

- can be an element
- elements pure substances, can not be broken down further (physical and chemical wise). ex. gold
- Have a freezing and boiling point that do not change
- Compounds are another example of a pure substance. Compounds are a chemical combination of two or more elements. Can be broken down into element by heat or electricity
- Example of a Compound: $\text{H}_2\text{O} = 2$ hydrogen molecules + one oxygen molecule

Mix Substance:

- Matter that are not compounds or elements can be classified as either a homogenous mixture (solution) or heterogeneous mixture (mechanical mixture). These mixtures contain two or more pure substances.
- Particles of mixtures look and act different from each other.
- Mixtures have boiling, freezing and melting points that are not set.

Homogenous Mixture (Solution):

- Homo = the same
- A homogenous mixture (solution) is a physical combination of two or more chemicals mixed evenly. The solution looks the same
- Is often clear. They cannot be filtered to separate the particles. If you shine a light through, it will not scatter the light.
- Example: Water and sugar (when the sugar is fully dissolved), PEPSI, brewed coffee, air etc.

Heterogeneous Mixture

- Hetero = different
- A heterogeneous mixture (mechanical mixture) is a physical combination of two or more chemicals that are not evenly mixed and that looks different all throughout.
- You can see the different substances that are combined together.
- Example: Chocolate chip cookie, salsa, salad dressing, pizza etc.

Study Guide Matter Of Change

Joe Ben Hoyle

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Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition McGraw Hill, 2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats. There are worksheets and study tools for each section of the text that help teachers track students' progress toward understanding concepts. Guided Reading Activities help students identify and comprehend the important information in each chapter.

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Schreck, 1993 **Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples **Shipman**

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