

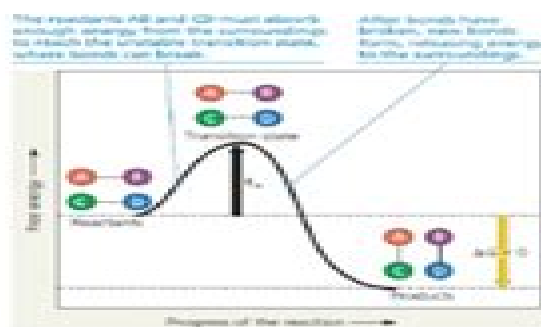
Study Guide: Enzymes

Enzymes are chemical agents that speed up a chemical reaction without being consumed. It usually ends with -ase.

A spontaneous chemical reaction occurs without any requirement for outside energy, but it may occur so slowly that it is imperceptible. For example, even though the hydrolysis of sucrose (table sugar) to glucose and fructose is exergonic, occurring spontaneously with a release of free energy ($G = -7$ kcal/mol), a solution of sucrose dissolved in sterile water will sit for years at room temperature with no appreciable hydrolysis. However, if we add a small amount of the enzyme sucrase to the solution, then all the sucrose may be hydrolyzed within seconds, as shown here:



How exactly do enzymes speed up the chemical reaction? Let's check the graph below.



- Chemical reaction involves breaking or forming a bond. e.g., sucrose \gg glucose and fructose
- Changing one molecule into another generally involves contorting the starting molecule into a highly unstable state before the reaction can proceed.
- Contorting the reactants is like trying to insert a key into the key holder. The keying is highly unstable in its opened form but returns to a stable state once the key is threaded all the way onto the ring.
- To reach the contorted, molecules need to absorb energy from their surroundings
- Energy required to contort the reactant molecules is called *Activation Energy*. It is supplied with thermal energy.
- The energy accelerates the reactant molecules and collides more often causing the breakage of a bond.



Study Guide On Enzymes

**Marcia L. Gillette, Ira Blei, Wendy
Gloffke, George Odian**



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Table of Contents Study Guide On Enzymes

1. Understanding the eBook Study Guide On Enzymes
 - The Rise of Digital Reading Study Guide On Enzymes
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide On Enzymes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide On Enzymes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide On Enzymes
 - Personalized Recommendations
 - Study Guide On Enzymes User Reviews and Ratings
 - Study Guide On Enzymes and Bestseller Lists
5. Accessing Study Guide On Enzymes Free and Paid eBooks

- Study Guide On Enzymes Public Domain eBooks
- Study Guide On Enzymes eBook Subscription Services
- Study Guide On Enzymes Budget-Friendly Options
- 6. Navigating Study Guide On Enzymes eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide On Enzymes Compatibility with Devices
 - Study Guide On Enzymes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide On Enzymes
 - Highlighting and Note-Taking Study Guide On Enzymes
 - Interactive Elements Study Guide On Enzymes
- 8. Staying Engaged with Study Guide On Enzymes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide On Enzymes
- 9. Balancing eBooks and Physical Books Study Guide On Enzymes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide On Enzymes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide On Enzymes
 - Setting Reading Goals Study Guide On Enzymes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide On Enzymes
 - Fact-Checking eBook Content of Study Guide On Enzymes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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