

SOLUTION CONCENTRATION UNITS

- **MASS %**: $\frac{\text{mass solute}}{\text{mass solution}} \times 100$
- **PPT**: Equation above $\times 1000$
- **PPM**: $\times 1,000,000$
- **PPB**: $\times 1,000,000,000$
- **MOLE FRACTION**: $X = \frac{\text{moles solute}}{\text{moles solution}}$
- **MOLARITY**: $M = \frac{\text{mol solute}}{\text{L solution}}$
- **MOLALITY**: $m = \frac{\text{moles solute}}{\text{kg solvent}}$

VAPOR PRESSURE OF A SOLUTION W/ NONVOLATILE

- **NONVOLATILE** = NO VP
- **ADDING NONVOLATILE SOLUTE TO SOLVENT LOWERS VP** B/C IT TAKES UP SPACE IN THE SOLUTION & INTERFERES W/ ESCAPE OF SOLVENT
- **RAOULT'S LAW** SHOWS RELATIONSHIP B/W SOLVENT VP & CONCENTRATION

$$P_1 = X \cdot P_1^0$$

VP solvent over solution mole fraction solvent VP pure solvent

FACTORS AFFECTING SOLUBILITY

- "LIKE DISSOLVES LIKE"
- **POLAR SOLVENTS TEND TO DISSOLVE POLAR MOLECULES & IONIC COMPOUNDS**
 - B/C POS/NEG END OF POLAR SOLVENT (LIKE H₂O) ATTRACTED TO POS/NEG END OF POLAR MOLECULE
- **NONPOLAR SOLVENTS DISSOLVE NONPOLAR MOLECULES**
- **FOR SOLIDS, SOLUBILITY ↑ AS TEMP ↑**
- **FOR GAS, SOLUBILITY ↓ AS TEMP ↑**
- **FOR LIQUID AND SOLID**:
PRESSURE HAS LITTLE EFFECT ON SOLUBILITY IN H₂O
- **FOR GAS**: HIGHER PARTIAL PRESSURE, THE MORE SOLUBLE
- **DEFINE BY HENRY'S LAW**:

$$C_g = K \cdot P_{\text{gas}}$$

concentration constant partial P.

VAPOR PRESSURE OF SOLUTION W/ VOLATILE

- **IN IDEAL SOLUTION WHERE SOLUTE AND SOLVENT FOLLOW RAOULT'S**, VP = SUM OF VP FROM EACH COMPONENT
- **DO RAOULT'S LAW TWICE**:
 $P_A = X_A \cdot P_A^0$
 $P_B = X_B \cdot P_B^0$
 $P_{\text{total}} = P_A + P_B$

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Brielle Morgan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Study Guide Solution Concentration:

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exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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Study Guide to Accompany Chemical Principles, Third Edition Daniel L. Reger, 1985

Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney William H. Myers, William T. Mooney, 1980

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