

SOLUTION CONCENTRATION UNITS

- MASS %: $\frac{\text{mass solute}}{\text{mass solution}} \times 100$

- PPT: Equation above $\times 1000$

- PPM: $\times 1,000,000$

- PPB: $\times 1,000,000,000$

- MOLE FRACTION: $X = \frac{\text{moles solute}}{\text{moles solution}}$

- MOLARITY: $M = \frac{\text{moles solute}}{\text{L solution}}$

- MOLALITY: $m = \frac{\text{moles solute}}{\text{kg solvent}}$

VAPOR PRESSURE OF A SOLUTION W/ NONVOLATILE

- NONVOLATILE = NO VP

- ADDING NONVOLATILE SOLUTE TO SOLVENT LOWERS VP B/C IT TAKES UP SPACE IN THE SOLUTION & INTERFERS W/ ESCAPE OF SOLVENT

- RAOUULT'S LAW SHOWS RELATIONSHIP B/W SOLVENT VP & CONCENTRATION

$$P_i = X_i \cdot P_i^0$$

VP Solvent over Solution \checkmark mole fraction Solvent \checkmark VP pure Solvent

FACTORS AFFECTING SOLUBILITY

- "LIKE DISSOLVES LIKE"

- POLAR SOLVENTS TEND TO DISSOLVE POLAR MOLECULES & IONIC CMPDS

- B/C POS/NEG END OF POLAR SOLVENT (LIKE H₂O) ATTRACTED TO POS/NEG END OF POLAR MOLECULE

- NONPOLAR SOLVENTS DISSOLVE NONPOLAR MOLECULES

- FOR SOLIDS, SOLUBILITY \uparrow AS TEMP \uparrow

- FOR GAS, SOLUBILITY \downarrow AS TEMP \uparrow

- FOR LIQUID AND SOLID: PRESSURE HAS LITTLE AFFECT ON SOLUBILITY IN H₂O

- FOR GAS: HIGHER PARTIAL PRESSURE, THE MORE SOLUBLE

- DEFINE BY HENRY'S LAW:

$$C_g = K \cdot P_{\text{gas}}$$

concentration \checkmark partial constant P.

VAPOR PRESSURE OF SOLUTION W/ VOLATILE

- IN IDEAL SOLUTION WHERE SOLUTE AND SOLVENT FOLLOW RAOUULT'S. VP = SUM OF VP FROM EACH COMPONENT

- DO RAOUULT'S LAW TWICE:

$$P_A = X_A \cdot P_A^0 \quad P_{\text{total}} = P_A + P_B$$

$$P_B = X_B \cdot P_B^0$$

Study Guide Solution Concentration

Brielle Morgan

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