

# Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

## Chapters 17 & 18

**Essential Skills/ State Standards:**

9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)

9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)

9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CSS: 7d)

9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)

9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d\*, 8c)

9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 8a)

9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

**Vocabulary:**

<ul style="list-style-type: none"> <li>• Condensation</li> <li>• Freezing</li> <li>• Vaporization</li> <li>• Melting</li> <li>• Boiling</li> <li>• Sublimation</li> </ul>	<ul style="list-style-type: none"> <li>• Deposition</li> <li>• Heat</li> <li>• Temperature</li> <li>• Thermochemistry</li> <li>• Exothermic</li> <li>• Endothermic</li> </ul>	<ul style="list-style-type: none"> <li>• Calorimeter</li> <li>• Joule</li> <li>• Specific heat</li> <li>• <math>q = m \cdot c \cdot \Delta T</math></li> <li>• <math>q = m \cdot \Delta H</math> <small>exothermic</small></li> <li>• <math>q = m \cdot \Delta H</math> <small>endothermic</small></li> </ul>	<ul style="list-style-type: none"> <li>• Heat of reaction</li> <li>• Le Chatelier's principle</li> <li>• Equilibrium</li> <li>• Activation Energy</li> <li>• Catalyst</li> </ul>
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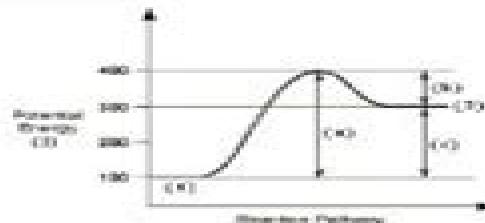
**Book Assignments:**

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Section 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

**Study Guide (Use notes, worksheets, and labs in addition to this study guide)**

1. According to the collision theory, what 3 things must happen in order for two molecules to react?
2. What is the importance of the activation energy in a chemical reaction?
3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:  

$$2 \text{ HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$
  - a. Increase temp:
  - b. Increase HCl
  - c. Increase pressure
  - d. Remove  $\text{MgCl}_2$
  - e. Add  $\text{H}_2$
4. Identify the following terms with their correct number/letter from the graphs:
  - \* Reactants- \_\_\_\_\_
  - \* Products- \_\_\_\_\_
  - \* Activation energy (number/letter & amount in Joules)- \_\_\_\_\_
  - \* Is this reaction exothermic or endothermic? \_\_\_\_\_
5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



# Study Guide Stoichiometry Content Mastery Answers

**Pamela Mills,Amina El-Ashmawy**

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a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text *The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Chemistry Study Guide* Gregory Gebhart,2017-09-21 Solutions to stoichiometry and conservation of mass problems Laboratory exercises

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**Chemistry Lesson Plans, Study Guides, and Lecture Notes** Doctor Chemistry,2017-08-13 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slide of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and Molar Mass Equation Balancing Thermodynamics Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Acids and Bases Limiting Reagents Redox and Electro Chemistry Organic Chemistry Basics **Selected Solutions for Chemical Principles** Kenneth C. Brooks,Steven S. Zumdahl,1992 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Study Guide and Selected Solutions Manual for Basic Chemistry** Karen C. Timberlake,2016-03-29 The Study Guide and Selected Solutions Manual assists students with the text material It contains learning objectives chapter outlines additional problems with self tests and answers and answers to the odd numbered problems in the text

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