

## Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

### Chapters 17 & 18

**Essential Skills/ State Standards:**

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)  
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)  
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CSS: 7d)  
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)  
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d\*, 8c)  
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)  
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

**Vocabulary**

- |   |   |  |  |
|---|---|--|--|
| <ul style="list-style-type: none"> <li>• Condensation</li> <li>• Freezing</li> <li>• Vaporization</li> <li>• Melting</li> <li>• Boiling</li> <li>• Sublimation</li> </ul> | <ul style="list-style-type: none"> <li>• Deposition</li> <li>• Heat</li> <li>• Temperature</li> <li>• Thermochemistry</li> <li>• Exothermic</li> <li>• Endothermic</li> </ul> | <ul style="list-style-type: none"> <li>• Calorimeter</li> <li>• Joule</li> <li>• Specific heat</li> <li>• <math>q = m \cdot c \cdot \Delta T</math></li> <li>• <math>q = m \cdot \Delta H_{\text{vaporization}}</math></li> <li>• <math>q = m \cdot \Delta H_{\text{fusion}}</math></li> </ul> | <ul style="list-style-type: none"> <li>• Heat of reaction</li> <li>• Le Chatelier's principle</li> <li>• Equilibrium</li> <li>• Activation Energy</li> <li>• Catalyst</li> </ul> |
|---|---|--|--|

**Book Assignments:**

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

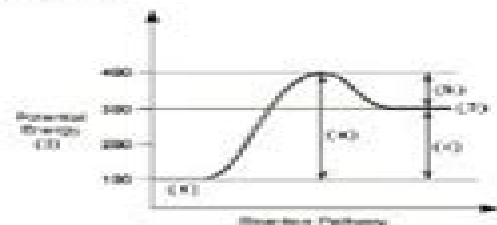
**Study Guide (Use notes, worksheets, and labs in addition to this study guide)**

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:  

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove $\text{MgCl}_2$	e. Add $\text{H}_2$
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- Identify the following terms with their correct number/letter from the graphs:
 

• Reactants-	• Products-
• Activation energy (number/letter & amount in Joules)-	
• Is this reaction exothermic or endothermic?	
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



# Study Guide Stoichiometry Content Mastery Answers

**Pamela Mills,Amina El-Ashmawy**



## Study Guide Stoichiometry Content Mastery Answers:

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