

Name: _____

Atomic Structure & the Periodic Table Study Guide

Atomic Structure

Particle	Charge	Where are they found?
Proton	Positive	Nucleus
Neutron	Neutral/ no charge	Nucleus
Electron	negative	Electron cloud/energy levels outside nucleus

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**

2. What are the vertical columns on the periodic table called? **Groups or families**

a. How many of these are on the periodic table? **18**

b. Why are the elements in these columns? **They have the same chemical properties (& same # of valence electrons)**

c. Elements in the same group have the same number of **valence** electrons.

3. What are the horizontal rows on the periodic table called? **periods**

a. How many of these are on the periodic table? **7**

b. Elements in these rows have the same number of **energy (electron shells)** levels.

4. What are valence electrons? **Electrons in the outer energy level**

5. How many electrons are in each energy level?

1st energy level **2** 2nd energy level **8** 3rd energy level **8**

6. The atomic number is equal to the **# of protons and also # of electrons**

7. The atomic mass is equal to the **# protons + # of neutrons**

8. How do you find the number of neutrons? **Atomic mass – atomic number**

9. We view the order of the periodic table today according to the **atomic number**

Study Guide The Periodic Table Answer Key

**William K. Purves, Edward
Dzialowski, Lindsay Goodloe, Betty
McGuire, Nancy Guild, Paula Mabee**

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Press, Raymond Siever, 2003-09-25 This reconceptualization of the text Understanding Earth reflects the fundamental changes in the field of physical geology over the past several years

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