

Thermodynamics Study Guide

Equations that you need to know:

$$PV = nRT = Nk_B T \quad (\text{ideal gas law})$$

$$\text{KE} = \frac{3}{2} k_B T \quad (\text{translational energy of gas molecules})$$

$$Q = mc \Delta T \quad (\text{def. of specific heat capacity})$$

$$Q = mL \quad (\text{def. of latent heat})$$

$$\Delta U = Q - W \quad (\text{first law / energy conservation})$$

$$W = P \Delta V = \text{area under } P\text{-}V \text{ graph} \quad (\text{work done during expansion})$$

$$\Delta S = \frac{Q}{T} \quad (\text{increase in entropy due to heat input})$$

$$\text{engine efficiency} = \frac{\text{benefit}}{\text{cost}} = \frac{W_{\text{net}}}{Q_h}$$

$$\text{refrigerator coef. of performance} = \frac{\text{benefit}}{\text{cost}} = \frac{Q_c}{W_{\text{net}}}$$

I'll give you the formulas for thermal expansion, heat conduction, and radiation:

$$\frac{\Delta L}{L_0} = \alpha \Delta T \quad \frac{\Delta V}{V_0} = \beta \Delta T \quad \frac{Q}{\Delta t} = \frac{k_t A (T_h - T_c)}{\Delta x} \quad \frac{Q}{\Delta t} = \epsilon \sigma A T^4$$

You should understand the following terms/concepts:

- temperature
- thermal equilibrium
- ideal gas
- mole
- rms speed
- phase transformation
- heat
- work
- conduction
- convection
- radiation
- isothermal
- adiabatic
- entropy
- second law of thermodynamics

Finally, you should know the approximate numerical values of the following: room temperature in kelvin; atmospheric pressure in N/m^2 ; the number of liters in a m^3 ; the gas constant; Avogadro's number; Boltzmann's constant; and the number of joules in a calorie (i.e., the specific heat capacity of water).

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The Editors of REA, Ralph Pike



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