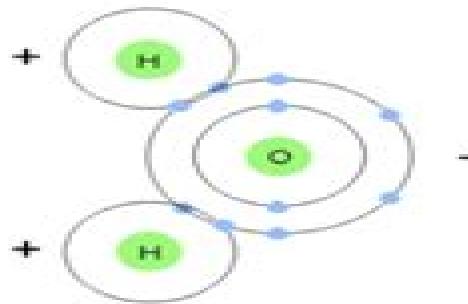


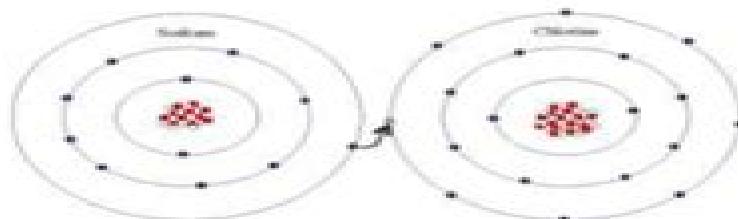
Chemical Bonding (PS. 4c)

Test date _____

Essential Concepts and Skills	Assignments
<p>The Formation of Compounds through Ionic and Covalent Bonding (PS. 4c)</p> <ul style="list-style-type: none"> Metals tend to lose electrons in chemical reactions, forming positive ions. Nonmetals tend to gain electrons in chemical reactions, forming negative ions. Given a chemical formula of a compound, identify the elements and the number of atoms of each that comprise the compound. Recognize that the number of electrons on the outermost energy level (valence electrons) determines an element's chemical properties or chemical reactivity. Describe the difference between ionic and covalent bonding. Predict what kind of bond (ionic or covalent) will likely form when metals and nonmetals are chemically bonded. Gaining or losing electrons makes an atom an ion. Atoms react to form chemically stable substances that are held together by chemical bonds and are represented by chemical formulas. To become chemically stable, atoms gain, lose, or share electrons. Compounds are formed when elements chemically react. When a metallic element reacts with a nonmetallic element, their atoms gain and lose electrons respectively, forming ionic bonds. Generally, when two nonmetals react, atoms share electrons, forming covalent (molecular) bonds. 	<p style="text-align: center;">Pages 328-345 Starter Quiz 1, 2 Gizmos</p>



Covalent Bond



Ionic Bond

Study Guide Ionic And Covalent Bonding

Paul B. Kelter

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