

Chapter 6:1

How did chemists begin to organize the known elements?

Early chemists used the properties of elements to sort them into groups.

How did Mendeleev organize his periodic table?

Mendeleev arranged the elements in his periodic table in order of increasing atomic mass.

How is the modern periodic table organized?

In the modern periodic table, elements are arranged in order of increasing atomic number.

What are the three broad classes of elements?

The three classes of elements are metals, nonmetals, and metalloids.

The properties of the elements within a period change as you move across the period from left to right. However, the pattern of properties within the period repeats as you move from one period to the next. This pattern gives rise to the periodic law: when elements are arranged in order of increasing atomic number, there is a periodic repetition of their physical and chemical properties. The arrangement of the elements into periods has an important consequence: elements that have similar chemical and physical properties end up in the same column in the periodic table.

Metals are generally good conductors of heat and electric current.

Nonmetals tend to have properties that are opposite to those of metals. In general, nonmetals are poor conductors of heat and electric current.

A metalloid generally has properties that are similar to those of metals and nonmetals. Under some conditions, a metalloid may behave like a metal. Under other conditions, it may behave like a nonmetal.

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