

4. a.
$$\frac{10 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33 \text{ g Cu}$$

$$\frac{20 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45 \text{ g Cu}}$$

b.
$$\frac{30 \text{ g Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25 \text{ g AlCl}_3$$

$$\frac{100 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11 \text{ g AlCl}_3}$$

c.
$$\frac{5 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67 \text{ g Cu}$$

~~100 g / 17.67 g~~ $10 \text{ g} / 17.67 \text{ g} \times 100 = \boxed{56.6\%}$

d.
$$\frac{10 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61 \text{ g AlCl}_3$$

~~100 g / 6.61 g~~ $5 \text{ g} / 6.61 \text{ g} \times 100 = \boxed{75.6\%}$

- 5.
- a. TRUE
 - b. FALSE
 - c. $\text{AgNO}_3 \Rightarrow 169.91 \text{ g/mol}$
 - d. 6.02×10^{23} particles = 1 mole of particles

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Table of Contents Study Guide Stoichiometry Answers

1. Understanding the eBook Study Guide Stoichiometry Answers
 - The Rise of Digital Reading Study Guide Stoichiometry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Stoichiometry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Stoichiometry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Stoichiometry Answers
 - Personalized Recommendations
 - Study Guide Stoichiometry Answers User Reviews and Ratings
 - Study Guide Stoichiometry Answers and Bestseller Lists
5. Accessing Study Guide Stoichiometry Answers Free and Paid eBooks
 - Study Guide Stoichiometry Answers Public Domain eBooks
 - Study Guide Stoichiometry Answers eBook Subscription Services
 - Study Guide Stoichiometry Answers Budget-Friendly Options
6. Navigating Study Guide Stoichiometry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Stoichiometry Answers Compatibility with Devices
 - Study Guide Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Stoichiometry Answers
 - Highlighting and Note-Taking Study Guide Stoichiometry Answers
 - Interactive Elements Study Guide Stoichiometry Answers
8. Staying Engaged with Study Guide Stoichiometry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Stoichiometry Answers
- 9. Balancing eBooks and Physical Books Study Guide Stoichiometry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Stoichiometry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Stoichiometry Answers
 - Setting Reading Goals Study Guide Stoichiometry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Stoichiometry Answers
 - Fact-Checking eBook Content of Study Guide Stoichiometry Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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