

## Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{100.0 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{56.0 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{26.9 \text{ g Al}} \cdot \frac{1 \text{ mol Al}}{1 \text{ mol Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{159.62 \text{ g Cu}}{1 \text{ mol Cu}} = 16.6 \text{ g Cu}$$

$$\frac{9.65 \text{ g CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{1 \text{ mol CuSO}_4}{1 \text{ mol CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{1 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.0\%}$$

3. If you start with 5.0 g of C<sub>3</sub>H<sub>8</sub> and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{2 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{3.22 \text{ g H}_2\text{O}}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}}$$

$$\text{PY} \cdot \text{TY} = \text{AY} = \frac{75 \cdot 3.22}{100} = \boxed{6.22 \text{ g H}_2\text{O}}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}}$$

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**Ralph Thomas Myers**

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## **Table of Contents Theoretical Yield Answer Key**

1. Understanding the eBook Theoretical Yield Answer Key
  - The Rise of Digital Reading Theoretical Yield Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Theoretical Yield Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Theoretical Yield Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Theoretical Yield Answer Key
  - Personalized Recommendations
  - Theoretical Yield Answer Key User Reviews and Ratings
  - Theoretical Yield Answer Key and Bestseller Lists
5. Accessing Theoretical Yield Answer Key Free and Paid eBooks
  - Theoretical Yield Answer Key Public Domain eBooks
  - Theoretical Yield Answer Key eBook Subscription Services

- Theoretical Yield Answer Key Budget-Friendly Options
- 6. Navigating Theoretical Yield Answer Key eBook Formats
  - ePUB, PDF, MOBI, and More
  - Theoretical Yield Answer Key Compatibility with Devices
  - Theoretical Yield Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Theoretical Yield Answer Key
  - Highlighting and Note-Taking Theoretical Yield Answer Key
  - Interactive Elements Theoretical Yield Answer Key
- 8. Staying Engaged with Theoretical Yield Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Theoretical Yield Answer Key
- 9. Balancing eBooks and Physical Books Theoretical Yield Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Theoretical Yield Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Theoretical Yield Answer Key
  - Setting Reading Goals Theoretical Yield Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Theoretical Yield Answer Key
  - Fact-Checking eBook Content of Theoretical Yield Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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