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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital m signifies solutions labelled with molar concentration a 1.0 M solution contains 1 mole of solute per litre of solution molality is the number of solvent moles per kilogram

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web 1. molarity mv mass molar mass $\times 0.4500$ l $\times 0.825$ g $\times 141.9579$ g/mol $\times 0.0129$ m $\times 2$ molality 0.825 g $\times 141.9579$ g/mol $\times 0.00581158$ mol $\times 0.00581158$ mol $\times 0.4500$ kg $\times 0.0129$ m $\times 3$ mole fraction n_2 hpo 4.0825 g $\times 141.9579$ g/mol $\times 0.00581158$ mol $\times 2$ o 450.0 g $\times 18.015$ g/mol $\times 24.97918401$ mol

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web apr 7 2014 molarity moles of solute liters of solution 8.422 a first convert 250 ml to liters $250/1000 = 0.25$ then calculate molarity 5 moles $/ 0.25$ liters $= 20$ M 3 c a solution with molarity 2 requires 2 M of NaOH per liter so 4×2.8 M $= 11.2$ M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula $m_{mix} = \frac{m_1 + m_2 + m_3}{V_1 + V_2 + V_3}$ $\frac{1.5 \times 480 + 1.2 \times 520 + 0.8 \times 720}{624 + 1000 + 1344}$ M q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1 calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 l of a solution 2 what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3 140 g of KCl is dissolved in 600 ml of water what is the molarity 4 724.4 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions 0.444 mol of CoCl_2 in 0.654 l of solution 98.0 g of phosphoric acid H_3PO_4 in 1.00 l of solution 0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 ml of solution 10.5 kg

16.11 molality chemistry libretexts - Aug 15 2023

web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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