

Stoichiometry**Answer Key**

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

(Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? **176 grams of sugar**
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? **15 sticks of butter**

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



Reactants

Fe_2O_3 (s) + 3CO (g) \rightarrow 2Fe (s) + 3CO₂ (g)

Reactants

How many moles of carbon monoxide (CO) are required to react completely with 1.25 moles of iron (III) oxide (Fe₂O₃)?

Products required: 0

1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe₂O₃: **1**

CO: **3**

Fe: **2**

CO₂: **3**

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants

Iron: **2**

Oxygen: **6**

Carbon: **3**

Products

Iron: **2**

Oxygen: **6**

Carbon: **3**

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Susan McMurry

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