

1.	<b>What is the third law of thermodynamics?</b>	The entropy of a perfect crystal is zero.
2.	<b>What is the second law of thermodynamics?</b>	As a whole, the universe always tends towards increasing entropy.
3.	<b>What is the relationship between number of microstates and entropy?</b>	More microstates = higher entropy
4.	<b>What is the relationship between temperature and entropy?</b>	Temperature increases, entropy increases (direct)
5.	<b>What is the relationship between atomic weight and entropy?</b>	Heavier = increased entropy (direct)
6.	<b>What happens to entropy when a substance is dissolved in another?</b>	Entropy increases
7.	<b>What is the relationship between moles of gas and entropy?</b>	When moles of gas increase, entropy increases (direct)
8.	<b>What is the relationship between free energy, enthalpy, and entropy?</b>	$\Delta G(\text{system}) = \Delta H(\text{system}) - T\Delta S(\text{system})$
9.	<b>What does change in G measure?</b>	The extent of the spontaneity of a process, and the useful energy available from it
10.	<b>What is the relationship between work and spontaneous processes?</b>	Spontaneous processes produce work
11.	<b>What is the relationship between work and non-spontaneous processes?</b>	Non-spontaneous processes require work

# Study Guide Thermochemistry

**James E. Brady, Fred Senese**



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