

## Ch. 5 Minerals – Quiz STUDY GUIDE

Name \_\_\_\_\_ Due Date \_\_\_\_\_ Block \_\_\_\_\_ Earth Science

1. Most minerals are compounds (contains atoms of two or more elements that are combined).
2. Native Elements (also called Native Minerals) contain only one element.
3. The internal arrangement of atoms determines the hardness, crystal shape and breakage of a mineral.
4. Minerals that break evenly along smooth, flat planes are said to have cleavage.
5. Luster refers to how light is reflected off the surface of a mineral.
6. Friedrich Mohs developed the hardness scale for minerals.
7. Minerals on the hardness scale are arranged according to their resistance to being scratched.
8. Color is an unreliable test for mineral identification.
9. Silicon, Oxygen, and Aluminum are the three most abundant elements in the Earth's crust.
10. A streak test is used to find the color of the powder of a mineral.
11. Scratching a mineral across a glass plate is a method used to determine a mineral's hardness.
12. What property of quartz accounts for its abundance? hardness
13. The mineral Quartz ( $\text{SiO}_4$ ) belongs to the silicates group.
14. The mineral Calcite fizzes when hydrochloric acid is placed on it.
15. The mineral Hematite is an ore of iron and has a characteristic reddish-brown streak.
16. The mineral Pyrite is also known as "Fool's Gold."
17. The mineral Mica breaks into thin, flat sheets.
18. The hardness of a fingernail is 2.5. Name two minerals, from the lab, that are softer than human fingernail: Talc & Gypsum.
19. Gemstones are located on Mohs scale at a hardness level of at least 7.
20. The yellow mineral sulfur is easily identified by its smell.
21. Which group of rock-forming minerals makes up 90% of Earth's crust? Silicates.

# Study Guide Minerals Ch 3

**T. W. Graham Solomons**

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physical world around us    **Organic Chemistry, Study Guide** T. W. Graham Solomons,1996 This supplement accompanies an updated text which features more than 350 new problems Carbonyl chemistry is now covered in consecutive chapters The concept of ionic reactions is consolidated before students move on to study radical reactions    **Study Guide** Stanley Chernicoff,2001-08 From Edward E Chatelain Valdosta State University Georgia this study guide helps students review and master the key ideas from every chapter through labeling exercises Chapter Reviews with matching statements plus Practice Tests and Challenge Tests that consist of multiple choice true false matching and short essay questions

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