

Single Replacement Reaction

A single replacement reaction occurs when one element replaces another element in a compound.



Examples:



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A single replacement reaction or single displacement reaction is when one element replaces another in a compound.

Get the definition of a single replacement reaction or single displacement reaction. Get examples of single replacement reactions and learn how to use the [metal reactivity series](#) to predict whether a reaction will occur and the products.

Single Replacement Reaction Definition

A **single replacement reaction** is a chemical reaction where one element replaces another in a compound. It is also known as a **single displacement reaction**. The general form of a single replacement reaction chemical equation is:



Single replacement reactions occur when A is more reactive than B or **product** AC is more stable than BC. A and B can be either two metals (including hydrogen; C is [an anion](#)) or else two halogens (C is a cation). If BC and AC are in aqueous solutions, C acts as a spectator ion.

Single Replacement Reaction Examples

There are two different scenarios for single replacement reactions. In one form of the reaction, one cation replaces the other. In the other form of the reaction, one anion replaces the other.

Cation Replacement Examples

Usually the cation is a metal, but it doesn't have to be. Here are examples of single replacement reactions involving the cations:

Single Replacement Reactions Study Guide

Leo J. Malone

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

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