

Equilibrium Practice Worksheet (Answers)

- The percent reaction of a reaction is 70%. Explain whether the equilibrium favours the reactants or products? What do you expect the magnitude of K to be? (Favours the products. Expect K value >> 0.)
- Write an equilibrium constant expression for the reactions below.
 - $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ ($K = (\text{NH}_3)^2 / (\text{N}_2)(\text{H}_2)^3$)
 - $\text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightleftharpoons \text{BaSO}_4(\text{aq}) + 2\text{NaNO}_3(\text{aq})$ ($K = (\text{NaNO}_3)^2 / (\text{Ba}(\text{NO}_3)_2)(\text{Na}_2\text{SO}_4)$)
- Determine the K_c for the reaction when 0.9 M of sulfur dioxide and 0.35 M of oxygen form 1.1 M of sulfur trioxide. ($K_c = 4.3$)
- Using the reaction from question 3. Determine which direction the equilibrium will shift if the Q value is 1×10^{-4} . (Will shift move toward the products)
- If the K_c value for an equilibrium is 3.8×10^{-6} , are the products or reactants favoured? What would the K_c value for the reverse reaction? (2.6×10^5)
- What would you expect the K value to be for the following graphs? (Would expect the first graph to have a K value of just above 1 and the second to have just below 1.)



- Determine which direction the equilibrium will shift when the following stresses are added.



- Temperature is increased - right
- Pressure is decreased - right
- Oxygen is added - left
- A catalyst is added - no shift
- Neon gas is added - no shift

- Graph the changes when for a, c and d. - (For a show increase in the products, decrease in reactants. For c should show an upward increase for O_2 and then a rate increase for reactants and rate decrease for products. For c should not see a change.)
- Which direction would you expect the equilibrium to shift if NaOH was added to the reaction below. (Would shift right)

Skill Practice 45 Equilibrium Practice Answers

**Laurence D. Hoffmann, Gerald L.
Bradley, Kenneth H. Rosen**

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