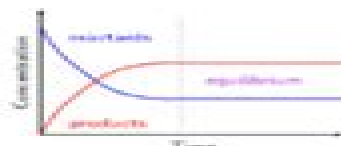


## Equilibrium Practice Worksheet (Answers)

- The percent reaction of a reaction is 70%. Explain whether the equilibrium favours the reactants or products? What do you expect the magnitude of K to be? (Favours the products. Expect K value  $\gg 0$ .)
- Write an equilibrium constant expression for the reactions below.
  - $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$  ( $K = (\text{NH}_3)^2 / (\text{N}_2)(\text{H}_2)^3$ )
  - $\text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightleftharpoons \text{BaSO}_4(\text{s}) + 2\text{NaNO}_3(\text{aq})$  ( $K = (\text{NaNO}_3)^2 / (\text{Ba}(\text{NO}_3)_2)(\text{Na}_2\text{SO}_4)$ )
- Determine the  $K_c$  for the reaction when 0.9 M of sulfur dioxide and 0.35 M of oxygen form 1.1 M of sulfur trioxide. ( $K_c = 4.3$ )
- Using the reaction from question 3. Determine which direction the equilibrium will shift if the Q value is  $1 \times 10^{-4}$ . (Will shift move toward the products)
- If the  $K_c$  value for an equilibrium is  $3.8 \times 10^{-6}$ , are the products or reactants favoured? What would the  $K_c$  value for the reverse reaction? ( $2.6 \times 10^5$ )
- What would you expect the K value to be for the following graphs? (Would expect the first graph to have a K value of just above 1 and the second to have just below 1.)



- Determine which direction the equilibrium will shift when the following stresses are added.



- Temperature is increased - **right**
  - Pressure is decreased - **right**
  - Oxygen is added - **left**
  - A catalyst is added - **no shift**
  - Neon gas is added - **no shift**
- Graph the changes when for a, c and d. - (For a show increase in the products, decrease in reactants. For c should show an upward increase for  $\text{O}_2$  and then a rate increase for reactants and rate decrease for products. For c should not see a change.)
  - Which direction would you expect the equilibrium to shift if NaOH was added to the reaction below. (Would shift right)

# Skill Practice 45 Equilibrium Practice Answers

**Laurence D. Hoffmann, Gerald L.  
Bradley, Kenneth H. Rosen**



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