

Name _____

Stoichiometry

Questions and Answers

1. What is the mole ratio in the equation $2H_2 + O_2 \rightarrow 2H_2O$?
2:1:2
2. Define stoichiometry.
Stoichiometry is the calculation of reactants and products in chemical reactions based on the balanced equation.
3. How do you determine the limiting reactant?
By calculating the amount of product formed from each reactant and identifying the reactant that forms the least product.
4. What is the theoretical yield?
The maximum amount of product that can be formed from given reactants.
5. How is percent yield calculated?
Percent Yield = (Actual Yield / Theoretical Yield) \times 100%
6. Given $N_2 + 3H_2 \rightarrow 2NH_3$, how many moles of NH_3 are produced from 3 moles of N_2 ?
6 moles
7. What are excess reactants?
Reactants that remain after the limiting reactant is completely used up.
8. Calculate the theoretical yield if the limiting reactant produces 2 moles of product.
2 moles (converted to grams if needed)
9. Define actual yield.
The amount of product actually obtained from the reaction.
10. If 5 moles of A react with 5 moles of B in the reaction $A + 2B \rightarrow AB_2$, which is the limiting reactant?
 B

Stoichiometry Review Study Guide Answer Key

James E. Brady



Stoichiometry Review Study Guide Answer Key:

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