

Name \_\_\_\_\_

# Stoichiometry

## Questions and Answers

1. What is the mole ratio in the equation  $2H_2 + O_2 \rightarrow 2H_2O$ ?

2:1:2

2. Define stoichiometry.

*Stoichiometry is the calculation of reactants and products in chemical reactions based on the balanced equation.*

3. How do you determine the limiting reactant?

*By calculating the amount of product formed from each reactant and identifying the reactant that forms the least product.*

4. What is the theoretical yield?

*The maximum amount of product that can be formed from given reactants.*

5. How is percent yield calculated?

*Percent Yield = (Actual Yield/Theoretical Yield) × 100%*

6. Given  $N_2 + 3H_2 \rightarrow 2NH_3$ , how many moles of  $NH_3$  are produced from 3 moles of  $N_2$ ?

6 moles

**PREVIEW**

7. What are excess reactants?

*Reactants that remain after the limiting reactant is completely used up.*

8. Calculate the theoretical yield if the limiting reactant produces 2 moles of product.

2 moles (converted to grams if needed)

9. Define actual yield.

*The amount of product actually obtained from the reaction.*

10. If 5 moles of A react with 5 moles of B in the reaction  $A + 2B \rightarrow AB_2$ , which is the limiting reactant?

B

# Stoichiometry Review Study Guide Answer Key

**James E. Brady**

## **Stoichiometry Review Study Guide Answer Key:**

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