

Stoichiometry Study Guide

Students should be able to...

- Explain mole ratio
- Convert moles, grams, atoms/molecules
- Calculate theoretical yields when given the limiting reactant (Convert: moles A to moles B, grams A to moles B, moles A to grams B, grams A to grams B)
- Explain limiting reactant and excess.
- Calculate theoretical yields by finding the limiting reactant
- Calculate percent yield.
- Calculate percent composition.
- Calculate theoretical yield, limiting reactant, excess, and percent yield for a double replacement reaction (laboratory)

Practice Problems

Find the molar mass of each compound.

1. BaI_2
2. $\text{AgC}_2\text{H}_3\text{O}_2$

Convert.

3. 0.4 moles of Sn to atoms
4. 32 g of K_2O to molecules
5. 3.7 g of H_2 to moles
6. 800 g of CaCl_2 to moles
7. 0.75 moles of H_2O to grams
8. 9.0 g of Cl_2 to atoms
9. 0.5 moles of KNO_3 to grams

Find the percent composition for each element in the compound.

10. Mn_2O_7
11. AuNO_3

Stoichiometry



If you started with 3.7 moles of magnesium oxide, how many moles of magnesium hydroxide would be produced?



If 2.8 grams of sodium were used for the reaction, and an excess of iodine was present, how many moles of sodium iodide would be produced?



Unit 7 Stoichiometry Study Guide

Robert K. Wismer

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Unit 7 Stoichiometry Study Guide:

Study Guide and Problems Book for Biochemistry, Garrett and Grisham David Karl Jemiolo, Garrett, Charles M. Grisham, 1996 *Student Study Guide to Accompany Petrucci's General Chemistry* Robert K. Wismer, 1985 *Resources in Education*, 1980 Robinson Chemistry Study Guide Robinson, 1992 Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993 *Research in Education*, 1974 Study Guide to Accompany Elements of Chemistry, General, Organic, and Biological -- Robert Boikess/Kenneth Breslauer/Edward Edelson Robert A. Klein, 1986

Structure, Bonding and the Periodic Law The Open University, Open University. Faculty of Science, Open University. Structure, Bonding and the Periodic Law Course Team, 1973 **Study Guide General Chemistry, Second Edition, Becker/Wentworth** Becker, Benjamin L. Carroll, 1980 **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 **Study Guide** Joan I. Senyk, Darrell D. Ebbing, 1984 **Chemistry, Study Guide** Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988 Chemistry, Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 CHEMISTRY SILBERBERG, 2003 Cooperative Chemistry Lab Manual Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a

problem There are no formalized instructions and students must plan and carry out their own investigations

McGraw-Hill Encyclopedia of Science and Technology ,1971

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