

# Unit 6: Molecules & Compounds

## Key Terms

Ionic Bonding  
Electron dot diagram  
ion  
anion  
cation  
chemical bond  
ionic bond  
chemical formula  
crystals

Covalent Bonding  
Covalent bond  
molecule  
polar covalent bond

Naming Compounds & Writing Formulas  
Polyatomic ion

## Essential Questions

Electron Configurations: **electron dot diagram** – a model of an atom in which each dot represents a valence electron.

PERIODIC TABLE ELEMENTS 1-20									
H									He
Li	Be	B	C	N	O	F			Ne
Na	Mg	Al	Si	P	S	Cl			Ar
K	Ca								

### • What is a chemical bond?

A **chemical bond** is the force that holds atoms and ions together as a unit.

### • How is an ionic bond different from a covalent bond?

**Ionic Bonds:** some elements achieve stable electron configurations through the transfer of electrons between atoms.



Formation of ions – when an atom gains or loses an electron, the number of protons is no longer equal to the number of electrons. The charge is no longer neutral.

**Ion** – an atom with a net positive or negative charge. Represented by a + or – sign.

Cl gains an electron 17 protons 18 electrons charge of 1- Cl<sup>1-</sup> or Cl<sup>-</sup>

**Anion** – (an eye un) an ion with a negative charge

**Naming:** using part of the element name plus the suffix -ide.

Cl<sup>-</sup>: Chloride ion.

Na loses an electron 11 protons 10 electrons charge of 1+ Na<sup>1+</sup> or Na<sup>+</sup>

**Cation** – (kat eye un) an ion with a positive charge

**Naming:** just use the elements name.

# Unit 6 Study Guide Covalent Bonding

**Martha R. Taylor**

## **Unit 6 Study Guide Covalent Bonding:**

**Ssg- Human Biology 6E Student Study Guide** Chiras,2008-02 Human Biology Sixth Edition provides students with a clear and concise introduction to the general concepts of mammalian biology and human structure and function. With its unique focus on health and homeostasis Human Biology enhances students understanding of their own health needs and presents the scientific background necessary for students to think critically about biological information they encounter in the media. The completely revised content and exceptional new art and photos provide students with a more user friendly text while excellent learning tools maximize comprehension of material [Study Guide to Accompany Basics for Chemistry](#)

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Mcdougal Littell,2004 [Life: The Science of Biology Study Guide](#) William K. Purves,Edward Dzialowski,Lindsay Goodloe,Betty McGuire,Nancy Guild,Paula Mabee,2003-12-26 New edition of a text presenting underlying concepts and showing their relevance to medical agricultural and environmental issues. Seven chapters discuss the cell information and heredity evolutionary process the evolution of diversity the biology of flowering plants and of animals and ecology and biogeography. Topics are linked by themes such as evolution the experimental foundations of knowledge the flow of energy in the living world the application and influence of molecular techniques and human health considerations. Includes a CD ROM

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Martha B. Barrett,1986 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today**

Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Student Guide For Living Chemistry* Barbara Ucko,2012-12-02 Student Guide for Living Chemistry is a 23 chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit Each objective corresponds to a section in the textbook Living Chemistry Three kinds of questions are provided for each objective to check the student's understanding namely short answer Study Questions multiple choice and fill in The answers for all questions are provided at the end of the chapter The opening chapters cover the SI units composition of matter chemical bonding compounds chemical change gases respiration and water The subsequent chapters deal with solutions acids bases salts nuclear and organic chemistry oxygen derivatives and hydrocarbons polymers and other organic derivatives This textbook also explores the chemistry of carbohydrates lipids proteins enzymes and energy and carbohydrate metabolism The remaining chapters discuss the chemistry of vitamins hormones body fluid drugs and poisons Undergraduate chemistry students will find this book invaluable

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