

## Chemistry Unit 5 Study Guide:

**Significant Figure:** each of the digits of a number that are used to express it to the required degree of accuracy, starting from nonzero digit.

### **Rules for determining sig figs:**

- All non-zero numbers are significant
- Zeros between non-zero digits are significant ex 7003
- Trailing zeros (those at the end) are only significant if the number contains a decimal point. Ex:  
1,000.0=5 sig figs 7,000=1 sig fig
- Zeros to the left of the first non-zero are insignificant, they serve as place holder. Ex: 00.01

### **Calculations with sig figs:**

#### Multiplication/Division:

Your answer should have the same amount of sig figs as the measurement with the fewest significant figures.

Ex:  $3.021\text{m} \times 9.0000\text{m} \times 8.2\text{m} = 222.9498\text{m}^3 \rightarrow 220$  because only has 2 sig figs.

#### Addition/subtraction:

Your answer should have the same number of digits to the right of the decimal as the measurement having the fewest number of digits to the right decimal.

Ex:  $5.001\text{m} + 98.5\text{m} + 5001.2315 = 5104.7325 \rightarrow 5104.7$

**Scientific Notation:** a way for scientists to handle a very large or very small number in a more organized manner. Needs a coefficient (1-9.99), base, and exponent.

Negative exp- small number- move to the left

Positive exp- big number- move to the right

**Moles:** in chemistry they are  $6.02 \times 10^{23}$  particles (atoms, molecules, or formula units) of a substance (elements and compounds)

1 mole =  $6.02 \times 10^{23}$

### Converting between molecules and particles:

- Figure out which units you are converting between
- Set up a proportion
- Look at what the problem gives you. Multiply this by a conversion factor formed from your equality in step 2.

Ex: how many moles of Mg is  $1.25 \times 10^{23}$  atoms of magnesium.

a) Atoms-moles

$1/6.02 \times 10^{23} = x/1.25 \times 10^{23}$  x has to be the same amount of sig fig as the number in the problem.

### Converting Between moles and mass:

**Grand Formula Mass/Molar Mass:** the mass of one mole of a substance. Expressed in grams.

- Calculate GFM
- Create conversion factor.



# Unit 5 Study Guide Chemistry Answers

**Darrell D. Ebbing**



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