

**Multiple Choice Section:** This study guide is a compilation of questions from provincial exams since April 1994. I urge you to become intimately familiar with question types. You will notice that questions from one year to another are very similar in their composition. Identification of question types will allow you to be more efficient in answering these questions on the provincial examination. My recommendations for using this study guide are as follows:

1. **DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet!
2. **RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
3. **LEARN FROM YOUR MISTAKES!** If you get a question wrong, **figure out why!** If you are having difficulty, **talk to your study partner**, or maybe **phone someone in your Peer Tutoring group**. Get together with group members or other students from class and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it!) Ask your teacher to explain the questions to you during tutorial or after school. **Your goal should be to get 100% on any Chemistry 12 multiple choice test**—learning from your mistakes in this booklet will really help you in your efforts to meet this goal!
4. **This is REALLY CRUCIAL: DO NOT mark the answer anywhere on the questions themselves.** For example, do not circle any of options A B C or D—instead use a different sheet of paper to place your answers on. By avoiding this urge, you can re-use this study guide effectively again, when preparing for your final exam. In the box to the left, put an asterisk or small note to yourself to indicate that you got the question wrong and need to come back to it. If you got the question correct initially, a check mark might be assurance that you understand this type of question and therefore can concentrate on other questions that present a challenge to you.
5. **Check Off the STATUS box on the PRESCRIBED LEARNING OUTCOMES sheet.** I have tried to organize the questions in the identical sequence to which they appear on your Acid Base Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the UNIT EXAM and PROVINCIAL EXAM!

INTRODUCTION TO REACTION KINETICS.....	1
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1.	A2	The rate of a chemical reaction can be expressed in A. grams per mole. C. volume of gas per unit time. B. energy consumed per mole. D. moles formed per litre of solution.
2.	A2	Consider the following reaction: <b>Under certain conditions, the rate of decomposition of <math>\text{NO}_2</math> is <math>3.2 \times 10^{-3} \text{ mol/s}</math>. The rate of formation of <math>\text{O}_2</math> is:</b> A. $1.6 \times 10^{-3} \text{ mol/s}$ B. $3.2 \times 10^{-3} \text{ mol/s}$ C. $4.8 \times 10^{-3} \text{ mol/s}$ D. $6.4 \times 10^{-3} \text{ mol/s}$
3.	A2	Consider the following reaction: $2\text{N}_2\text{O}_{5(g)} \rightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$ <b>At a certain temperature the rate of decomposition of <math>\text{N}_2\text{O}_5</math> is <math>2.5 \times 10^{-6} \text{ mol/s}</math>. The rate of formation of <math>\text{NO}_2</math> is:</b> A. $1.0 \times 10^{-6} \text{ mol/s}$ B. $1.3 \times 10^{-6} \text{ mol/s}$ C. $2.5 \times 10^{-6} \text{ mol/s}$ D. $5.0 \times 10^{-6} \text{ mol/s}$
4.	A2	Consider the following reaction: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ <b>The rate of formation of <math>\text{NH}_3</math> is <math>3.0 \text{ mL/min}</math>. The rate of consumption of <math>\text{H}_2</math> is</b> A. $1.5 \text{ mL/min}$ B. $2.0 \text{ mL/min}$ C. $4.5 \text{ mL/min}$ D. $9.0 \text{ mL/min}$
5.	A2	Which of the following units can be used to express the rate of a chemical reaction? A. $\text{mL/g}$ B. $\text{mol/L}$ C. $\text{g/mol}$ D. $\text{mol/min}$

# Study Guide Chemical Kinetics

**Kenneth Antonio Connors**



## Study Guide Chemical Kinetics:

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