

**THE OFFICIAL STUDY GUIDE FOR:  
"CHEMICAL KINETICS"**

**Multiple Choice Section:** This study guide is a compilation of questions from provincial exams since April 1994. I urge you to become intimately familiar with question types. You will notice that questions from one year to another are very similar in their composition. Identification of question types will allow you to be more efficient in answering these questions on the provincial examination. My recommendations for using this study guide are as follows:

1. **DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet!
  2. **RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
  3. **LEARN FROM YOUR MISTAKES!** If you get a question wrong, **figure out why!** If you are having difficulty, **TALK to your study partner**, or maybe **phone someone to post Peer Tutoring access**. Get together with group members or other students from class and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it)! Ask your teacher to explain the questions to you during tutorial or after school. **Your goal should be to get 100% on any Chemistry 12 multiple choice test**—learning from your mistakes in this booklet will really help you in your efforts to meet this goal!
  4. **This is REALLY CRUCIAL: DO NOT mark the answer anywhere on the questions themselves.** For example, do not circle any of options A, B, C, or D—instead use a different sheet of paper to place your answers on. By avoiding this urge, you can re-use this study guide effectively again, when preparing for your final exam. In the box to the left, put an asterisk or small note to yourself to indicate that you got the question wrong and need to come back to it. If you got the question correct initially, a check mark might be assurance that you understand this type of question and therefore can concentrate on other questions that present a challenge to you.
  5. **Check Off the STATUS box on the PRESCRIBED LEARNING OUTCOMES sheet.** I have tried to organize the questions in the identical sequence to which they appear on your Acid Base Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the **UNIT EXAM** and **PROVINCIAL EXAM**!

## TABLE OF CONTENTS

INTRODUCTION TO REACTION KINETICS.....	1
COLLISION THEORY.....	6
REACTION MECHANISMS AND CATALYSTS.....	11
ANSWERS.....	21

## INTRODUCTION

1. A2 The rate of a chemical reaction can be expressed in  
A. grams per mole. B. energy consumed per mole.  
C. volume of gas per unit time. D. moles formed per litre of solution.

2. A2 Consider the following reaction:  
Under certain conditions, the rate of decomposition of  $\text{NO}_2$  is  $3.2 \times 10^{-3}$  mol/s. The rate of formation of  $\text{O}_2$  is:  
A.  $1.6 \times 10^{-3}$  mol/s  
B.  $3.2 \times 10^{-3}$  mol/s  
C.  $4.8 \times 10^{-3}$  mol/s  
D.  $6.4 \times 10^{-3}$  mol/s

3. A2 Consider the following reaction:

$$2\text{N}_2\text{O}_5(\text{g}) \rightarrow 4\text{NO}_{2(\text{g})} + \text{O}_{2(\text{g})}$$

At a certain temperature the rate of decomposition of  $\text{N}_2\text{O}_5$  is  $2.5 \times 10^{-6}$  mol/s. The rate of formation of  $\text{NO}_2$  is:  
A.  $1.0 \times 10^{-5}$  mol/s B.  $1.3 \times 10^{-6}$  mol/s C.  $2.5 \times 10^{-6}$  mol/s D.  $5.0 \times 10^{-6}$  mol/s

4. A2 Consider the following reaction:



At a certain temperature the rate of decomposition of  $\text{N}_2\text{O}_5$  is  $2.5 \times 10^{-6}$  mol/s. The rate of formation of  $\text{NO}_2$  is:

4. A2 Consider the following reaction:



The rate of formation of  $\text{NH}_3$  is 3.0  $\text{mL/min}$ . The rate of consumption of  $\text{H}_2$  is

5. A2 Which of the following units can be used to express the rate of a chemical reaction?  
A. mL/g    B. mol/L    C. g/mol    D. mol/msin

# Study Guide Chemical Kinetics

**Kenneth Antonio Connors**

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*Study Guide, Chemical Principles, Fifth Edition, Zumdahl* Paul B. Kelter, Steven S. Zumdahl, 2004-04 **Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993 Chemical Principles Peter

Atkins, Loretta Jones, 2010 This text is designed for a rigorous course in introductory chemistry. Its central theme is to challenge students to think and question while providing a sound foundation in the principles of chemistry. Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem-solving strategies. Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts. Includes many illustrative worked examples.

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Resources in Education ,1978    Laboratory Directions and Study Questions in Inorganic Chemistry Alexander Silverman,1919    **Chemical Kinetics** Kenneth Antonio Connors,1990 Chemical Kinetics The Study of Reaction Rates in Solution Kenneth A Connors This chemical kinetics book blends physical theory phenomenology and empiricism to provide a guide to the experimental practice and interpretation of reaction kinetics in solution It is suitable for courses in chemical kinetics at the graduate and advanced undergraduate levels This book will appeal to students in physical organic chemistry physical inorganic chemistry biophysical chemistry biochemistry pharmaceutical chemistry and water chemistry all fields concerned with the rates of chemical reactions in the solution phase

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## **Table of Contents Study Guide Chemical Kinetics**

1. Understanding the eBook Study Guide Chemical Kinetics
  - The Rise of Digital Reading Study Guide Chemical Kinetics
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Chemical Kinetics
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Chemical Kinetics
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Chemical Kinetics
  - Personalized Recommendations

- Study Guide Chemical Kinetics User Reviews and Ratings
  - Study Guide Chemical Kinetics and Bestseller Lists
5. Accessing Study Guide Chemical Kinetics Free and Paid eBooks
- Study Guide Chemical Kinetics Public Domain eBooks
  - Study Guide Chemical Kinetics eBook Subscription Services
  - Study Guide Chemical Kinetics Budget-Friendly Options
6. Navigating Study Guide Chemical Kinetics eBook Formats
- ePUB, PDF, MOBI, and More
  - Study Guide Chemical Kinetics Compatibility with Devices
  - Study Guide Chemical Kinetics Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Study Guide Chemical Kinetics
  - Highlighting and Note-Taking Study Guide Chemical Kinetics
  - Interactive Elements Study Guide Chemical Kinetics
8. Staying Engaged with Study Guide Chemical Kinetics
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Chemical Kinetics
9. Balancing eBooks and Physical Books Study Guide Chemical Kinetics
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Chemical Kinetics
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Chemical Kinetics
- Setting Reading Goals Study Guide Chemical Kinetics
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Chemical Kinetics
- Fact-Checking eBook Content of Study Guide Chemical Kinetics

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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