

**CHAPTER 7 REVIEW*****Chemical Formulas and Chemical Compounds*****SECTION 7-1****SHORT ANSWER** Answer the following questions in the space provided.

1. c In a Stock name such as iron(III) sulfate, the roman numeral tells us \_\_\_\_\_.  
 (a) how many atoms of Fe are in one formula unit  
 (b) how many sulfate ions can be attached to the iron atom  
 (c) the charge on each Fe ion  
 (d) the total positive charge of the formula unit

2. c The result of changing a subscript in a correctly written chemical formula is to \_\_\_\_\_.  
 (a) change the number of moles represented by the formula  
 (b) change the charges on the other ions in the compound  
 (c) change the formula so that it no longer represents the compound it previously represented  
 (d) have no effect on the formula

3. The explosive TNT has the molecular formula  $\text{C}_7\text{H}_5(\text{NO}_2)_3$ .  
4 elements a. How many elements make up this compound?  
6 oxygen atoms b. How many oxygen atoms are present in one molecule of  $\text{C}_7\text{H}_5(\text{NO}_2)_3$ ?  
21 atoms c. How many atoms in total are present in one molecule of  $\text{C}_7\text{H}_5(\text{NO}_2)_3$ ?  
 $4.2 \times 10^{24}$  atoms d. How many atoms are present in a sample of  $2 \times 10^{23}$  molecules of  $\text{C}_7\text{H}_5(\text{NO}_2)_3$ ?

4. How many atoms are present in each of these formula units?  
11 atoms a.  $\text{Ca}(\text{HCO}_3)_2$   
45 atoms b.  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$   
10 atoms c.  $\text{Fe}(\text{ClO}_3)_3$   
9 atoms d.  $\text{Fe}(\text{ClO}_3)_2$

5.  $\text{N}_2\text{O}_5$  a. What is the formula for the compound dinitrogen pentoxide?  
carbon(IV) sulfide b. What is the Stock name for the covalent compound  $\text{CS}_2$ ?  
 $\text{H}_2\text{SO}_4$  c. What is the formula for sulfuric acid?  
phosphoric acid d. What is the name for the acid  $\text{H}_3\text{PO}_4$ ?

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**Paul A. Bartlett, Judith G. Koch**

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