

Chapter 16 - "Solutions" Worksheet Answer Key

1. Indicate whether the following are ionic compounds or covalent compounds:

a) NaCl <i>ionic</i>	b) CaBr ₂ <i>ionic</i>	c) SO ₂ <i>covalent</i>	d) NH ₃ <i>covalent</i>
e) H ₂ O <i>covalent</i>	f) Sr(OH) ₂ <i>ionic</i>	g) MgCO ₃ <i>ionic</i>	h) AgOH <i>ionic</i>
i) ZnCl ₂ <i>ionic</i>	j) SiO ₂ <i>covalent/ionic</i>		

2. Show the dissociation of the following substances:

a) KCl (s) \rightarrow K^+ (aq) + Cl^- (aq)

b) MgF₂ (s) \rightarrow Mg^{2+} (aq) + $2F^-$ (aq)

c) FeS (s) \rightarrow Fe^{2+} (aq) + S^{2-} (aq)

d) Hg₂SO₄ (s) \rightarrow $2Hg^{2+}$ (aq) + SO_4^{2-} (aq)

e) Ca₃(PO₄)₂ (s) \rightarrow $3Ca^{2+}$ (aq) + $2PO_4^{3-}$ (aq)

f) CCl₄ (l) \rightarrow *No dissociation (non-polar)*

g) Ag₂S (s) \rightarrow $2Ag^+$ (aq) + S^{2-} (aq)

h) NH₄Cl (s) \rightarrow NH_4^+ (aq) + Cl^- (aq)

i) (NH₄)₂Cr₂O₇ (s) \rightarrow $2NH_4^+$ (aq) + $Cr_2O_7^{2-}$ (aq)

j) K₂CrO₄ (s) \rightarrow $2K^+$ (aq) + CrO_4^{2-} (aq)

3. Write the net ionic equation for the dissociation of:

a) ammonium sulphate in water



b) iron (III) carbonate in water



4. Calculate the number of moles of all aqueous ions in the following solutions, assuming each

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Jane M. Cram,Donald J. Cram

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