

Chapter 16 - "Solutions" Worksheet Answer Key

1. Indicate whether the following are ionic compounds or covalent compounds:

- | | | | |
|-------------------------------------|---|------------------------------------|------------------------------------|
| a) NaCl <i>ionic</i> | b) CaBr ₂ <i>ionic</i> | c) SO ₂ <i>covalent</i> | d) NH ₃ <i>covalent</i> |
| e) H ₂ O <i>covalent</i> | f) Sr(OH) ₂ <i>ionic</i> | g) MgCO ₃ <i>ionic</i> | h) AgOH <i>ionic</i> |
| i) ZnCl ₂ <i>ionic</i> | j) SiO ₂ <i>covalent/ionic</i> | | |

2. Show the dissociation of the following substances:

- | | | | | |
|---|---|------------------------------------|---|---------------------|
| a) KCl (s) | → | $K^+ (aq)$ | + | $Cl^- (aq)$ |
| b) MgF ₂ (s) | → | $Mg^{2+} (aq)$ | + | $2F^- (aq)$ |
| c) FeS (s) | → | $Fe^{2+} (aq)$ | + | $S^{2-} (aq)$ |
| d) Hg ₂ SO ₄ (s) | → | $2Hg^+ (aq)$ | + | $SO_4^{2-} (aq)$ |
| e) Ca ₃ (PO ₄) ₂ (s) | → | $3Ca^{2+} (aq)$ | + | $2PO_4^{3-} (aq)$ |
| f) CCl ₄ (l) | → | <i>No dissociation (non-polar)</i> | | |
| g) Ag ₂ S (s) | → | $2Ag^+ (aq)$ | + | $S^{2-} (aq)$ |
| h) NH ₄ Cl (s) | → | $NH_4^+ (aq)$ | + | $Cl^- (aq)$ |
| i) (NH ₄) ₂ Cr ₂ O ₇ (s) | → | $2NH_4^+ (aq)$ | + | $Cr_2O_7^{2-} (aq)$ |
| j) K ₂ CrO ₄ (s) | → | $2K^+ (aq)$ | + | $CrO_4^{2-} (aq)$ |

3. Write the net ionic equation for the dissociation of:

- a) ammonium sulphate in water



- b) iron (III) carbonate in water



4. Calculate the number of moles of all aqueous ions in the following solutions, assuming each

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Jane M. Cram, Donald J. Cram



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