

**Stoichiometry****Answer Key**

**Vocabulary:** Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

*(Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.)*

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? **176 grams of sugar**
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? **15 sticks of butter**

**Gizmo Warm-up**

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



**Reactants**

$\text{Fe}_2\text{O}_3$  (s) + 3CO (g)  $\rightarrow$  2Fe (s) + 3CO<sub>2</sub> (g)

**Reactants**

How many moles of carbon monoxide (CO) are required to react completely with 1.25 moles of iron (III) oxide (Fe<sub>2</sub>O<sub>3</sub>)?

Products required: 0

1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe<sub>2</sub>O<sub>3</sub>: **1**

CO: **3**

Fe: **2**

CO<sub>2</sub>: **3**

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

**Reactants**

Iron: **2**

Oxygen: **6**

Carbon: **3**

**Products**

Iron: **2**

Oxygen: **6**

Carbon: **3**

# Study Guide Content Mastery Stoichiometry Answer Key

**Seager/Slabaugh**

## **Study Guide Content Mastery Stoichiometry Answer Key:**

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      **Resources in Education**, 1969

**Research in Education**, 1969      **Science Books & Films**, 1980      STOICHIOMETRY NARAYAN

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### **Stoichiometry Review**

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book's handy Periodic Table of Elements allows for quick answers on the elements appearing on the exam TABLE OF CONTENTS About Research and Education Association Independent Study Schedule CHAPTER 1 ABOUT THE SAT II CHEMISTRY SUBJECT TEST About This Book About The Test How To Use This Book Format of the SAT II Chemistry Scoring the SAT II Chemistry Score Conversion Table Studying for the SAT II Chemistry Test Taking Tips CHAPTER 2 COURSE REVIEW Gases Gas Laws Gas Mixtures and Other Physical Properties of Gases Dalton's Law of Partial Pressures Avogadro's Law The Mole Concept Avogadro's Hypothesis Chemical Compounds and Formulas Mole Concept Molecular Weight and Formula Weight Equivalent Weight Chemical Composition Stoichiometry Weight and Volume Calculations Balancing Chemical Equations Calculations Based on Chemical Equations Limiting Reactant Calculations Solids Phase Diagram Phase Equilibrium Properties of Liquids Density Colligative Properties of Solutions Raoult's Law and Vapor Pressure Osmotic Pressure Solution Chemistry Concentration Units Equilibrium The Law of Mass Action Kinetics and Equilibrium Le Chatelier's Principle and Chemical Equilibrium Acid Base Equilibria Definitions of Acids and Bases Ionization of Water pH Dissociation of Weak Electrolytes Dissociation of Polyprotic Acids Buffers Hydrolysis Thermodynamics I Bond Energies Some Commonly Used Terms in Thermodynamics The First Law of Thermodynamics Enthalpy Hess's Law of Heat Summation Standard States Heat of Vaporization and Heat of Fusion Thermodynamics II Entropy The Second Law of Thermodynamics Standard Entropies and Free Energies Electrochemistry Oxidation and Reduction Electrolytic Cells Non Standard State Cell Potentials Atomic Theory Atomic Weight Types of Bonds Periodic Trends Electronegativity Quantum Chemistry Basic Electron Charges Components of Atomic Structure The Wave Mechanical Model Subshells and Electron Configuration Double and Triple Bonds Organic Chemistry Nomenclature and Structure Alkanes Alkenes Dienes Alkynes Alkyl Halides Cyclic Hydrocarbons Aromatic Hydrocarbons Aryl Halides Ethers and Epoxides Alcohols and Glycols Carboxylic Acids Carboxylic Acid Derivatives Esters Amides Arenes Aldehydes and Ketones Amines Phenols and Quinones Structural Isomerism SIX PRACTICE EXAMS Practice Test 1 Answer Key Detailed Explanations of Answers Practice Test 2 Answer Key Detailed Explanations of Answers Practice Test 3 Answer Key Detailed Explanations of Answers Practice Test 4 Answer Key Detailed Explanations of Answers Practice Test 5 Answer Key Detailed Explanations of Answers Practice Test 6 Answer Key Detailed Explanations of Answers THE PERIODIC TABLE EXCERPT About Research OR 2 To demonstrate proficiency in Chemistry The SAT II Chemistry exam is designed for students who have taken one year of college preparatory chemistry Who Administers The Test The SAT II Chemistry Subject Test is developed by the College Board and administered by Educational Testing Service ETS The test development process involves the assistance of educators throughout the country and is designed and implemented to ensure that the content and difficulty level of the test are appropriate When Should the SAT II Chemistry be Taken If you are applying to a college that requires Subject Test scores as part of the admissions process you should take the SAT II Chemistry Subject Test toward the end of your junior year or at the beginning of your senior year If your scores are being

used only for placement purposes you may be able to take the test in the spring of your senior year For more information be sure to contact the colleges to which you are applying When and Where is the Test Given The SAT II Chemistry Subject Test is administered five times a year at many locations throughout the country mostly high schools To receive information on upcoming administrations of the exam consult the publication Taking the SAT II Subject Tests which may be obtained from your guidance counselor or by contacting College Board SAT Program P O Box 6200 Princeton NJ 08541 6200 Phone 609 771 7600 Website <http://www.collegeboard.com> Is There a Registration Fee Yes There is a registration fee to take the SAT II Chemistry Consult the publication Taking the SAT II Subject Tests for information on the fee structure Financial assistance may be granted in certain situations To find out if you qualify and to register for assistance contact your academic advisor

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**FORMAT OF THE SAT II CHEMISTRY** The SAT II Chemistry is a one hour exam consisting of 85 multiple choice questions The first part of the exam consists of classification questions This question type presents a list of statements or questions that you must match up with a group of choices lettered A through E Each choice may be used once more than once or not at all The exam then shifts to relationship analysis questions which you will answer in a specially numbered section of your answer sheet You will have to determine if each of two statements is true or false and if the second statement is a correct explanation of the first The last section is composed strictly of multiple choice questions with choices lettered A through E

**Material Tested** The following chart summarizes the distribution of topics covered on the SAT II Chemistry Subject Test

Topic	Percentage	Number of Questions
Atomic Structure	15%	12
Chemical Bonding	15%	12
Chemical Reactions	15%	12
Electrochemistry	10%	8
Periodic Table	10%	8
Properties of Matter	10%	8
States of Matter	10%	8
Thermodynamics	10%	8
Acids and Bases	5%	4
Organic Chemistry	5%	4
Other Topics	5%	4

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Atomic Structure 15% 12

Chemical Bonding 15% 12

Chemical Reactions 15% 12

Electrochemistry 10% 8

Periodic Table 10% 8

Properties of Matter 10% 8

States of Matter 10% 8

Thermodynamics 10% 8

Acids and Bases 5% 4

Organic Chemistry 5% 4

Other Topics 5% 4

however do not review too much at any one time Concentrate on one problem area at a time by reviewing the questions and explanations and by studying our review until you are confident you completely understand the material Keep track of your scores By doing so you will be able to gauge your progress and discover general weaknesses in particular sections You should carefully study the reviews that cover your areas of difficulty as this will build your skills in those areas

**TEST TAKING**

**TIPS** Although you may be unfamiliar with standardized tests such as the SAT II Chemistry Subject Test there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties. Become comfortable with the format of the exam. When you are practicing to take the SAT II Chemistry Subject Test simulate the conditions under which you will be taking the actual test. Stay calm and pace yourself. After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual exam with much more confidence. Know the directions and format for each section of the test. Familiarizing yourself with the directions and format of the exam will not only save you time but will also ensure that you are familiar enough with the SAT II Chemistry Subject Test to avoid nervousness and the mistakes caused by being nervous. Do your scratchwork in the margins of the test booklet. You will not be given scrap paper during the exam and you may not perform scratchwork on your answer sheet. Space is provided in your test booklet to do any necessary work or draw diagrams. If you are unsure of an answer guess. However if you do guess guess wisely. Use the process of elimination by going through each answer to a question and ruling out as many of the answer choices as possible. By eliminating three answer choices you give yourself a fifty fifty chance of answering correctly since there will only be two choices left from which to make your guess. Mark your answers in the appropriate spaces on the answer sheet. Fill in the oval that corresponds to your answer darkly completely and neatly. You can change your answer but remember to completely erase your old answer. Any stray lines or unnecessary marks may cause the machine to score your answer incorrectly. When you have finished working on a section you may want to go back and check to make sure your answers correspond to the correct questions. Marking one answer in the wrong space will throw off the rest of your test whether it is graded by machine or by hand. You don't have to answer every question. You are not penalized if you do not answer every question. The only penalty results from answering a question incorrectly. Try to use the guessing strategy but if you are truly stumped by a question remember that you do not have to answer it. Work quickly and steadily. You have a limited amount of time to work on each section so you need to work quickly and steadily. Avoid focusing on one problem for too long. Before the Test Make sure you know where your test center is well in advance of your test day so you do not get lost on the day of the test. On the night before the test gather together the materials you will need the next day. Your admission ticket. Two forms of identification e.g. driver's license, student identification card or current alien registration card. Two No 2 pencils with erasers. Directions to the test center. A watch if you wish but not one that makes noise as it may disturb other test takers. On the day of the test you should wake up early after a good night's rest and have breakfast. Dress comfortably so that you are not distracted by being too hot or too cold while taking the test. Also plan to arrive at the test center early. This will allow you to collect your thoughts and relax before the test and will also spare you the stress of being late. If you arrive after the test begins you will not be admitted to the test center and you will not receive a refund. During the Test When you arrive at the test center try to find a seat where you feel most comfortable. Follow all the rules and instructions given by the

test supervisor If you do not you risk being dismissed from the test and having your scores canceled Once all the test materials are passed out the test instructor will give you directions for filling out your answer sheet Fill this sheet out carefully since this information will appear on your score report After the Test When you have completed the SAT II Chemistry Subject Test you may hand in your test materials and leave Then go home and relax When Will I Receive My Score Report and What Will It Look Like You should receive your score report about five weeks after you take the test This report will include your scores percentile ranks and interpretive information Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition Leo J. Malone, 2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone's Basic Concepts of Chemistry **Principles of Modern Chemistry** Wade A. Freeman, David W. Oxtoby, 1996 PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course The fifth edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today Authors David W Oxtoby and H P Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process from observation to application placing general chemistry into a complete perspective for serious minded science and engineering students Chemical principles are illustrated by the use of modern materials comparable to equipment found in the scientific industry Students are therefore exposed to chemistry and its applications beyond the classroom This text is perfect for those instructors who are looking for a more advanced general chemistry textbook *Almond Books ICSE Class 10 CHEMISTRY Study Guide 2022-23 Notes + Question Bank + Sample Papers 2023 Exam (Fully Solved)* Almond Books, Chemistry is the key to understanding the world around us With Almond Books exclusive Chemistry study guide ICSE Class 10 students can master the periodic table chemical bonding metallurgy mole concept and more Get ready to ace your exams with the help of this must have resource The book follows the latest ICSE Chemistry syllabus and exam pattern provided in the most recent ICSE Chemistry specimen paper Contents of this book Chapter wise REVISION NOTES AND QUESTION BANK This study guide is packed with chapter wise revision notes important chemical equations and tips from board examiners Plus it's loaded with even more valuable content to help you ace your ICSE Chemistry exams The Fully solved question bank covers all essential exam content in one book It includes Multiple choice questions New and important questions short and long answers Popular ICSE school's prelim questions Important questions from previous year board exams SAMPLE PAPERS OFFICIAL SPECIMEN ANSWERS Chemistry Sample Papers 2 Question Papers with QR code solutions Answer key for latest official ICSE Chemistry Specimen Paper Chapters covered in this book Based on Concise Selina Evergreen Publication Textbooks Periodic Table Periodic Properties and Variations of Properties Chemical Bonding Acids Bases and Salts Analytical Chemistry Uses of Ammonium Hydroxide and Sodium Hydroxide Mole concept and Stoichiometry Electrolysis Metallurgy Study of Compounds Hydrogen Chloride Study of Compounds Ammonia Study of Compounds Nitric Acid Study of Compounds

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