

Name _____

Electrons in Atoms



Study Guide: Properties of Waves and Quantum Mechanics

Properties of Waves

- **Wavelength (λ):** The distance between successive crests of a wave, usually measured in meters.
- **Frequency (ν):** The number of waves that pass a given point per second, measured in Hertz (Hz).
- **Amplitude:** The height of the wave crest or depth of the trough, related to the wave's energy.
- **Speed of light (c):** The rate at which the wave propagates through a medium, calculated using $c = \lambda \nu$.

Calculating the Wavelength of a Wave

$$\lambda = c / \nu$$

- λ is the wavelength.
- c is the wave speed.
- ν is the frequency.

Speed of Light

- The speed of light in a vacuum (c) is approximately 3.00×10^8 meters per second (m/s).

The Quantum Equation

- Relates energy of a photon to its frequency: $E = h\nu$
 - E is the energy.
 - h is Planck's constant (6.626×10^{-34} Js).
 - ν is the frequency.

Element Emission Spectrum

The concept of quantized energy levels within atoms is crucial to understanding atomic spectra. When an electron absorbs energy, it jumps to a higher energy level, and when it returns to a lower energy level, it emits a photon of light with a specific energy corresponding to the difference in energy levels. This explains why each element has a unique spectral fingerprint, a set of specific wavelengths of light it emits or absorbs.

- When elements are heated, they emit light at specific wavelengths, creating an emission spectrum. This spectrum is unique to each element and can be used to identify them.



Study Guide Electrons In Atoms

Douglas Brandt



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