

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

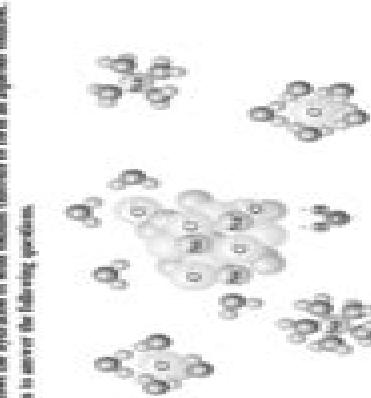
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



1. Indicate in solute is when the solute is in a (1) _____.

(2) _____ is the process of surrounding solute particles with solvent particles to form a (3) _____.

Study Guide for Content Mastery

Section 15.2 Solubility**Section 15.3 Solubility****STUDY GUIDE FOR CONSTITUTANT MASTERY****Section 15.1 content**

15. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are (1) _____ separated and surrounded by the water molecules.

16. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is (2) _____ positively charged, it attracts the (3) _____ negatively charged portion of the water molecule (the oxygen atom) and repels the (4) _____ positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is (5) _____ negatively charged, it attracts the (6) _____ positively charged portion of the water molecule and repels the (7) _____ negatively charged portion.

17. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the (8) _____ water molecules and the sodium and chloride ions is (9) _____ greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the (10) _____ solution process to occur.

18. In your textbook, read about heat of solution, solubility, and factors that affect solubility. For each statement below, write true or false.

True _____ 15. The overall energy change that occurs when a solute dissolves in a solvent is called the heat of solution.

False _____ 16. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False _____ 17. Solubility increases as long as the solute has a high heat of solution.

True _____ 18. In a saturated solution, solute and solvent molecules are in equilibrium.

True _____ 19. Additional solute can be dissolved in an unsaturated solution.

True _____ 20. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

Section 15.2 Solubility**Section 15.3 Solubility****Study Guide for Content Mastery****Study Guide for Content Mastery**

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**Thomas A. Stetz, Leslie A.
Miller, Robert L. Lovler**

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1. Understanding the eBook Study Guide For Content
 - The Rise of Digital Reading Study Guide For Content
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content
 - Personalized Recommendations
 - Study Guide For Content User Reviews and Ratings
 - Study Guide For Content and Bestseller Lists
5. Accessing Study Guide For Content Free and Paid eBooks
 - Study Guide For Content Public Domain eBooks
 - Study Guide For Content eBook Subscription Services
 - Study Guide For Content Budget-Friendly Options
6. Navigating Study Guide For Content eBook Formats

- ePub, PDF, MOBI, and More
- Study Guide For Content Compatibility with Devices
- Study Guide For Content Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Study Guide For Content
- Highlighting and Note-Taking Study Guide For Content
- Interactive Elements Study Guide For Content

8. Staying Engaged with Study Guide For Content

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide For Content

9. Balancing eBooks and Physical Books Study Guide For Content

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Study Guide For Content

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Study Guide For Content

- Setting Reading Goals Study Guide For Content
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide For Content

- Fact-Checking eBook Content of Study Guide For Content
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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